

Stoichiometry Lab - **ANSWERS****PRE-LAB QUESTIONS:**

1. Calculate the molar mass of sodium bicarbonate (NaHCO_3). 84.003 g/mol
 2. Calculate the molar mass of acetic acid (CH_3COOH). 60.052 g/mol
 3. Calculate the molar mass of carbon dioxide (CO_2). 44.003 g/mol
 4. Vinegar is generally 4-5% acetic acid dissolved in distilled water. That means that a 5% vinegar solution contains 0.05 mol of acetic acid per liter of solution. How many moles of acetic acid are present in 30.00 mL?
- $$\frac{0.05 \text{ mol CH}_3\text{COOH}}{1 \text{ liter vinegar}} \times \frac{0.03 \text{ liter CH}_3\text{COOH}}{1 \text{ liter vinegar}} = 0.0015 \text{ mol CH}_3\text{COOH in 30.0 mL of vinegar}$$
5. When baking soda (sodium bicarbonate) reacts with vinegar (acetic acid), they form sodium acetate, water, and carbon dioxide gas. Balance the chemical equation for this reaction.



6. A student measures out 5.00 g of baking soda to react with 50.00 mL of vinegar (which contains 0.042 moles of acetic acid).

a) Determine the limiting reagent, excess reagent, and theoretical yield (in grams) of carbon dioxide. Show your work.

$$5.00 \text{ g baking soda} \times \frac{1 \text{ mol baking soda}}{84.003 \text{ g baking soda}} = \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol baking soda}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 2.62 \text{ g carbon dioxide}$$

$$0.042 \text{ mole acetic acid} \times \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol acetic acid}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 1.85 \text{ g carbon dioxide}$$

** Since acetic acid produced the least amount of product, the acetic acid in vinegar is the **limiting reagent** and the baking soda is the **excess reagent**.

** The theoretical yield of carbon dioxide is 1.85 g.

- b) This reaction will produce carbon dioxide bubbles that escape. After the reaction, the mass of the reaction decreased by 1.50 g. This is the actual mass of carbon dioxide that was produced from the reaction. Calculate the percentage yield of carbon dioxide. Show your work.

$$\text{Percent Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent Yield} = \frac{1.50 \text{ g}}{1.85 \text{ g}} \times 100 = 79.4\% \text{ yield}$$

Stoichiometry Lab Vinegar And Baking Soda Answers

Priti Singhal



Stoichiometry Lab Vinegar And Baking Soda Answers:

A Concrete Stoichiometry Unit for High School Chemistry Jennifer Louise Pakkala, 2006 *A Stoichiometry Unit* David Callaghan, 2004 **CBSE Chemistry Chapterwise Case Study Class 11** Priti Singhal, 2024-11-17 This book is structured to align with the latest syllabus and curriculum guidelines ensuring that the content is both relevant and rigorous Each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter We have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging *Holt Chemistry* Ralph Thomas Myers, 2004 Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007 *Stoichiometry* Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay s Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904

Stoichiometry Unit Project Luann Marie Decker, 1998 Reacting Vinegar with Baking Soda Henry D. Schreiber, 1997-02-01 Formulation and Stoichiometry Emil J. Margolis, 1968 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **You Can Do Chemistry** Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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