

Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

- 3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Stoichiometry Chapter 12 Test A Answers

John A. Olmsted, Gregory M. Williams



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Table of Contents Stoichiometry Chapter 12 Test A Answers

1. Understanding the eBook Stoichiometry Chapter 12 Test A Answers
 - The Rise of Digital Reading Stoichiometry Chapter 12 Test A Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Chapter 12 Test A Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Chapter 12 Test A Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Chapter 12 Test A Answers
 - Personalized Recommendations
 - Stoichiometry Chapter 12 Test A Answers User Reviews and Ratings
 - Stoichiometry Chapter 12 Test A Answers and Bestseller Lists
5. Accessing Stoichiometry Chapter 12 Test A Answers Free and Paid eBooks
 - Stoichiometry Chapter 12 Test A Answers Public Domain eBooks
 - Stoichiometry Chapter 12 Test A Answers eBook Subscription Services
 - Stoichiometry Chapter 12 Test A Answers Budget-Friendly Options

6. Navigating Stoichiometry Chapter 12 Test A Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Chapter 12 Test A Answers Compatibility with Devices
 - Stoichiometry Chapter 12 Test A Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Chapter 12 Test A Answers
 - Highlighting and Note-Taking Stoichiometry Chapter 12 Test A Answers
 - Interactive Elements Stoichiometry Chapter 12 Test A Answers
8. Staying Engaged with Stoichiometry Chapter 12 Test A Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Chapter 12 Test A Answers
9. Balancing eBooks and Physical Books Stoichiometry Chapter 12 Test A Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Chapter 12 Test A Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Chapter 12 Test A Answers
 - Setting Reading Goals Stoichiometry Chapter 12 Test A Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Chapter 12 Test A Answers
 - Fact-Checking eBook Content of Stoichiometry Chapter 12 Test A Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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