



Name: Chelsey Williams Date: 3/31/22

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? 15 sticks of butter

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 1

CO 3

Fe 2

CO_2 3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe_2O_3)?

Problems solved: 0



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Table of Contents Stoichiometry Assessment Answers

1. Understanding the eBook Stoichiometry Assessment Answers
 - The Rise of Digital Reading Stoichiometry Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Assessment Answers
 - Personalized Recommendations
 - Stoichiometry Assessment Answers User Reviews and Ratings
 - Stoichiometry Assessment Answers and Bestseller Lists
5. Accessing Stoichiometry Assessment Answers Free and Paid eBooks

- Stoichiometry Assessment Answers Public Domain eBooks
- Stoichiometry Assessment Answers eBook Subscription Services
- Stoichiometry Assessment Answers Budget-Friendly Options
- 6. Navigating Stoichiometry Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Assessment Answers Compatibility with Devices
 - Stoichiometry Assessment Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Assessment Answers
 - Highlighting and Note-Taking Stoichiometry Assessment Answers
 - Interactive Elements Stoichiometry Assessment Answers
- 8. Staying Engaged with Stoichiometry Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Assessment Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Assessment Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Assessment Answers
 - Setting Reading Goals Stoichiometry Assessment Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Assessment Answers
 - Fact-Checking eBook Content of Stoichiometry Assessment Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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