

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before $\underline{2} \text{ mol}$ $\underline{0} \text{ mol}$ $\underline{0} \text{ mol}$

Change $\underline{-2 \text{ mol}}$ $\underline{+2 \text{ mol}}$ _____

After $\underline{0} \text{ mol}$ $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before $\underline{9.94} \text{ mol}$ $\underline{\text{xs mol}}$ $\underline{0} \text{ mol}$

Change $\underline{-9.94 \text{ mol}}$ _____ $\underline{+9.94 \text{ mol}}$

After $\underline{0} \text{ mol}$ $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before $\underline{1.0} \text{ mol}$ $\underline{\text{xs mol}}$ $\underline{0} \text{ mol}$ $\underline{0} \text{ mol}$

Change $\underline{-1.0 \text{ mol}}$ _____ $\underline{-6.0 \text{ mol}}$

After $\underline{0} \text{ mol}$ $\underline{\text{xs mol}}$

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

Stoichiometry Answer Key Simulation

Leonard Soltzberg



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Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs

Biosimulation Daniel A. Beard, 2012-04-12 This practical guide to biosimulation provides the hands on experience needed to devise design and analyze simulations of biophysical processes for applications in biological and biomedical sciences Through real world case studies and worked examples students will develop and apply basic operations through to advanced concepts covering a wide range of biophysical topics including chemical kinetics and thermodynamics transport phenomena and cellular electrophysiology Each chapter is built around case studies in a given application area with simulations of real biological systems developed to analyze and interpret data Open ended project based exercises are provided at the end of each chapter and with all data and computer codes available online www.cambridge.org/biosim students can quickly and easily run manipulate explore and expand on the examples inside This hands on guide is ideal for use on senior undergraduate graduate courses and also as a self study guide for anyone who needs to develop computational models of biological systems

Control and Safety Analysis of Intensified Chemical Processes Dipesh Shikchand Patle, Gade Pandu Rangaiah, 2024-03-25 Resource on the control and safety analysis of intensified chemical processes ranging from general methods to specific applications Control and Safety Analysis of Intensified Chemical Processes covers the basic principles of and recent developments in control and safety analysis of intensified chemical processes ranging from dynamic simulations and safety analysis to the design and control of important processes The text discusses general methods and tools such as dynamic simulation control and safety analysis as well as design aspects and analysis of important applications in order to provide scientists and engineers with an understanding of the design control and safety considerations involved in intensified chemical processes Sample topics covered in Control and Safety Analysis of Intensified Chemical Processes include Simulation and optimization methods common programs and simulators for simulation and optimization and interfacing of simulators and optimizers Programs simulators for dynamic simulation and control tuning of controllers and popular criteria for control assessment Control of a hybrid reactive extractive distillation systems for ternary azeotropic mixtures reactive distillation in recycle systems and middle vessel batch distillation with vapor recompression Safety analysis of intensified processes e g extractive distillation dividing wall column dividing wall column with mechanical vapor recompression and algal biodiesel process A comprehensive resource on the subject Control and Safety Analysis of Intensified Chemical Processes is a highly valuable reference for researchers students and practitioners

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