

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before **2 mol** **0 mol** **0 mol**

Change **-2 mol** **+2 mol**

After **0 mol** **2 mol**

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before **9.94 mol** **xs mol** **0 mol**

Change **-9.94 mol** **+9.94 mol**

After **0 mol** **9.94 mol**

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to

Stoichiometry 5 Percent Yield Answers

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Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests **STOICHIOMETRY**

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