Stoichiometry Review

1. The combustion of propone occurs via the reaction:

$$C_3H_3(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$$

How many grams of axygen are required to burn completely 10.0 g of propane?

2. Small quantities of chlorine can be prepared in the laboratory by the reaction:

What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

3. The most common one of arsenic is mispickel, FeSAs. Upon heating this one, free arsenic is obtained. How many grams of FeSAs are required to produce 10.0 g of arsenic according to the following reaction?

4. Glucose is used as an energy source by the human body. The overall reaction in the body is:

$$C_{i}H_{ij}O_{i}(aq) + 6 O_{i}(q) + 6 CO_{i}(q) + 5 H_{i}O(l)$$

Calculate the valume of oxygen at STP required to convert 28.0 g of glucose to carbon dioxide and water.

 Potassium nitrate is widely used as a fertilizer because it provides two essential elements, potassium and nitragen. It is made by mixing potassium chloride and nitric acid in the presence of axygen according to the equation

How many kilograms of patassium nitrate will be produced from 50.0 kg of patassium chloride and 50.0 kg of nitric acid with excess oxygen?

- 6. Bromine can be prepared by adding chlorine to an aqueous solution of sadium bromide. How many grams of bromine are formed if 25.0 g of chlorine and 25.0 g of sodium bromide are reacted according to the reaction below? Which reactant is in excess and how much remains after the reaction is complete?
 2 NoBr(ag) → Cl₂(g) → Br₂(I) → 2 NoCl(ag)
- How many millilitres of 2.00 M hydrochloric acid are required to react with 2.55 g of zinc according to the following reaction:

- B. A 0.473 g sample of phosphorus is reacted with an excess of chlorine. Calculate the mass of phosphorus pentachloride produced.
- 9. How many grams of exygen can be prepared by the decomposition of 25.0 g of mercury(III) exide?
- 10. What volume of oxygen gas at STP can be prepared by the decomposition of 25.0 g of potassium chlorate?
- 11. How many grams of zinc are required for the replacement of 0.100 g of hydrogen from sulfuric acid?
- 12. What mass of magnesium chloride is required to react with 10.0 g of silver nitrate?

Stoichiometry Section 2 Answers

Joseph M. Ranish, Charles W. Struck

Stoichiometry Section 2 Answers:

Regents Exams and Answers: Chemistry-Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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Oxidoreduction at the Plasma Membranerelation to Growth and Transport Frederick L. Crane, D. James Morre, Hans E. Low, 1991-04-15 The concept of general oxidoreductase function at the plasma level is new Oxidoreduction at the Plasma Membrane Relation to Growth and Transport provides the first truly comprehensive coverage of the oxidoreduction reactions in plasma membranes and the role that can now be attributed to these enzymes in controlling growth and other cell functions in plants and animals The book describes the nature and orientation of oxidoreductases in plasma membranes the stimulation of cell growth by oxidants reacting with transplasma membrane electron transport changes in enzymes in tumor cells and the basis of the growth effects and oxidoreductase stimulation of membrane transport in relation to known second messenger functions such as cellular pH changes calcium transport protein phosphorylation and oxidation levels of pyridine nucleotides The book then examines the significance of these enzymes in cell culture tumor growth nerve transmission ion transport and membrane potential The book also presents a new approach to understanding the action of antitumor drugs and herbicides Biochemists biologists oncologists cancer researchers physiologists and endocrinologists will find this an indispensible reference source as they conduct studies in this exciting new area Studies of High-Temperature Superconductors - Advances in Research and Applications Anant Narlikar, 2001 Studies of High Roadmap to the Regents Sasha Alcott, 2003 If Students Temperature Superconductors Volume 36 The BSCCO System II

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Inorganic Species, Part 1 Roger Minear, 2012-12-02 Inorganic Species Part 1 separately considers the various inorganic and organic components that occur in water While this separation is traditional it does provide some distinct organizational advantages This is important because of the wide ranging audience likely to be using these works Both practicing professionals and students in environmentally related disciplines will find these volumes to be a useful reference source This book comprises six chapters and begins with a focus on the origin and nature of selected inorganic constituents in natural waters Succeeding chapters go on to discuss redox potential which discusses its measurement and importance in water systems alkalinity and acidity conductance which is defined here as a collective measure of dissolved ions the theory and measurement of turbidity and residue and finally a summary of methods for water quality analysis of specific species. This book will be of interest to practitioners in the fields of geology and environmental engineering

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