

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

Stoichiometry Problems Year 11

L Reisser



Stoichiometry Problems Year 11:

The book delves into Stoichiometry Problems Year 11. Stoichiometry Problems Year 11 is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Stoichiometry Problems Year 11, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Stoichiometry Problems Year 11
- Chapter 2: Essential Elements of Stoichiometry Problems Year 11
- Chapter 3: Stoichiometry Problems Year 11 in Everyday Life
- Chapter 4: Stoichiometry Problems Year 11 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Stoichiometry Problems Year 11. The first chapter will explore what Stoichiometry Problems Year 11 is, why Stoichiometry Problems Year 11 is vital, and how to effectively learn about Stoichiometry Problems Year 11.
3. In chapter 2, this book will delve into the foundational concepts of Stoichiometry Problems Year 11. This chapter will elucidate the essential principles that need to be understood to grasp Stoichiometry Problems Year 11 in its entirety.
4. In chapter 3, the author will examine the practical applications of Stoichiometry Problems Year 11 in daily life. The third chapter will showcase real-world examples of how Stoichiometry Problems Year 11 can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Stoichiometry Problems Year 11 in specific contexts. This chapter will explore how Stoichiometry Problems Year 11 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Stoichiometry Problems Year 11. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Stoichiometry Problems Year 11.

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