

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

Stoichiometry Problems Worksheet Answer Key

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Levels Hala Raji Barakat,1998 Most of the strategies identified were either algorithmic correct strategies or incorrect strategies with incorrect answers where conceptual misunderstanding impeded successful problem solving The more conceptual thinkers could use the concepts to solve problems successfully in the test and argue correctly in the interview Moreover the learning approach was found not to be related either to conceptual understanding or to problem solving

strategies A major educational implication was that instruction should give more emphasis to conceptual understanding to enhance meaningful problem solving Also exams should include both the conceptual and the algorithmic aspects of problem solving

An Investigation Into Student Difficulties in Advanced Level Chemistry Problems Involving Stoichiometry Brian R. Jolliffe,1983 **Calculating in Chemistry** Cosmas C Mbuzi,2022-09-20 This book is intended to help students fully grasp

calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

Whispering the Secrets of Language: An Mental Journey through **Stoichiometry Problems Worsheet Answer Key**

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sklearn import tree dtree tree decisiontreeregressor min samples split 20 dtree fit x train y train print accuracy dtree predict
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