

Stoichiometry with Percent Yield

4) In manufacturing steel, iron oxide reacts with carbon to form iron and carbon monoxide as described by the reaction below. What is the percent yield of iron if 1400.3 grams of iron is produced from the reaction of 4.128 kg grams of Fe_2O_3 and excess carbon? The reaction is $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow 2\text{Fe} + \text{CO}$

5) 0.00001 grams of Barium-Chloride react with excess Sodium Sulfate. What is the percentage yield of Barium Sulfate if 0.0001280 grams of BaSO_4 is produced? $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$

6) 0.45009 grams of iron (II) sulfide reacts with nitrogen in the reaction below. If 0.58262 grams formed in the reaction, what is the percent yield? $2\text{Fe}_3\text{S}_4 + \text{N}_2 \rightarrow 4\text{Fe} + 2\text{N}_2\text{S}$

Name _____ Section _____ Date _____

Stoichiometry and Percent yield

Directions: Determine the amount of product that can be produced and the percent yield of each reaction.

1) 3.81 grams of H_3PO_4 react with excess NaOH in the following reaction: $\text{H}_3\text{PO}_4 + 3\text{NaOH} \rightarrow \text{Na}_3\text{PO}_4 + 3\text{H}_2\text{O}$. What is the percent yield if 3.05 grams of water are produced?

2) 9.199 grams of iron are produced when aluminum reacts with excess iron oxide in the reaction below. What is the percent yield if 3.95 grams of aluminum react? $\text{Fe}_2\text{O}_3 + \text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$

3) What is the percentage yield of CO_2 formed when 12.47 grams of CO_2 is collected after the reaction of 34.80 grams of C_2H_6 in the reaction: $\text{C}_2\text{H}_6 + \text{SO}_2 \rightarrow \text{SOO}_2 + \text{CO}_2$?

7) What is the percent yield when 5.502 grams of PbO_2 is produced in the reaction of 12.02 grams of Pb and excess $\text{Pb}(\text{NO}_3)_2$ in the reaction $\text{Pb}(\text{NO}_3)_2 + 2\text{Pb} \rightarrow \text{PbO}_2 + 2\text{PbNO}_2$

8) 21.28 grams of H_2S combine with excess oxygen. What is the percent yield if 14.48 grams of SO_2 are formed? $2\text{H}_2\text{S} + \text{SO}_2 \rightarrow 2\text{SO}_2 + 2\text{H}_2\text{O}$

9) When 13.3 grams of $\text{Ba}(\text{OH})_2$ reacts with excess $\text{HC}_2\text{H}_3\text{O}_2$, 6.08 grams of $\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$ are formed. What is the percent yield for this reaction? $\text{Ba}(\text{OH})_2 + 2\text{HC}_2\text{H}_3\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$

10) 95.800 grams of sulfur combine with excess oxygen and water to form 17.771 grams of sulfuric acid. What is the percent yield? In the reaction: $\text{S} + \text{SO}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SO}_4$

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Stoichiometry Percent Yield Answers

Neil D. Jespersen, Alison Hyslop

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Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career *Chemistry* Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order

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