

STOICHIOMETRY: MOLE-MOLE PROBLEMS



How many moles of hydrogen are needed to completely react with two moles of nitrogen?

$$2 \text{ mol } N_2 \left| \frac{3 \text{ mol } H_2}{1 \text{ mol } N_2} \right| = 6$$

6 moles H_2



How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

$$6 \text{ moles } KClO_3 \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} = \frac{18}{2} = 9$$

9 moles of O_2



How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?

$$3 \text{ mol } Zn \times \frac{1 \text{ mol } H_2}{1 \text{ mol } Zn} = 3$$

3 moles of H_2



How many moles of oxygen are necessary to react completely with four moles of propane (C_3H_8)?

$$4 \text{ mol } C_3H_8 \times \frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8}$$

20 mol of O_2



How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?

$$2 \text{ mol } K_3PO_4 \left| \frac{3 \text{ mol } KNO_3}{1 \text{ mol } K_3PO_4} \right| =$$

6 moles of KNO_3

Stoichiometry Packet Key

John T. Moore, Richard H. Langley



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or table salt 4 teaspoons cream of tarter 2 cups lukewarm water 2 tablespoons coconut oil or other oil coloring green 30
drops chlorophyll for 1 c dough yellow 1 teaspoon turmeric powder brighter pink 1 2

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