#### The Advanced Placement Examination in Chemistry

Part I – Multiple Choice Questions Part II - Free Response Questions Selected Questions from 1970 to 2010

#### Stoichiometry

#### Part I

1984

- Which of the following forms a compound having the formula KXO<sub>4</sub>?
  - (A) F
  - (B) S
  - (C) Mg.
  - (D) Ar (E) Mn
- 32. The net ionic equation for the reaction between silver carbonate and hydrochloric acid is
  - (A)  $Ag_2CO_3(s) + 2 H^* + 2 CT \rightarrow 2 AgCl(s) + H_2O + CO_2(g)$
  - (B)  $2 \text{ Ag}^{-} + \text{CO}_{3}^{2-} + 2 \text{ H}^{+} + 2 \text{ Cl}^{-} \rightarrow 2 \text{ AgCl}(s) + \text{H}_{2}\text{O} + \text{CO}_{2}(g)$
  - (C)  $CO_3^{2-} + 2 H^* \rightarrow H_2O + CO_2(g)$
  - (D)  $Ag' + CI' \rightarrow AgCl(s)$
  - (E)  $Ag_2CO_3(x) + 2 H^2 \rightarrow 2 Ag^2 + H_2CO_3$
- 44. What number of moles of O2 is needed to produce 14.2 grams of P4O10 from P? (Molecular weight
  - $P_aO_{10} = 284$ ) (A) 0.0500 mole
  - (B) 0.0625 mole
  - (C) 0.125 mole
  - (D) 0.250 mole
  - (E) 0.500 mole
- 45. The alkenes are compounds of carbon and hydrogen with the general formula C<sub>u</sub>H<sub>2v</sub>. If 0.561 gram of any alkene is burned in excess oxygen, what number of moles of H<sub>2</sub>O is formed?
  - (A) 0.0400 mole
  - (B) 0.0600 mole
  - (C) 0.0800 mole (D) 0.400 mole
  - (E) 0.800 mole
- 52.  $3 \text{ Ag}(s) + 4 \text{ HNO}_3 \rightarrow 3 \text{ AgNO}_3 + \text{NO}(g) + 2 \text{ H}_2\text{O}$

The reaction of silver metal and dilute nitric acid proceeds according to the equation above. If 0.10 mole of powdered silver is added to 10. milliliters of 6.0-molar nitric acid, the number of moles of NO gas that can be formed is

- (A) 0.015 mole
- (B) 0.020 mole
- (C) 0.030 mole
- (D) 0.045 mole
- (E) 0.090 mole

# **Stoichiometry Multiple Choice Questions**

**Almond Books** 

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