

Mole to Mole Stoichiometry Worksheet

- 1) In the reaction below, how many moles of aluminum will be used when reacted with 1.33 moles of oxygen?
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$

$$1.33 \text{ mol O}_2 \times \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = \boxed{1.77 \text{ mol Al}}$$

- 2) How many moles of hydrogen will be produced when reacted with 0.0204 moles of sodium in the reaction?
 $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + 1 \text{ H}_2$

$$0.0204 \text{ mol Na} \times \frac{1 \text{ mol H}_2}{2 \text{ mol Na}} = \boxed{0.0102 \text{ mol H}_2}$$

- 3) Using the equation $2 \text{ Al(NO}_3)_3 + 3 \text{ MgBr}_2 \rightarrow 3 \text{ Mg(NO}_3)_2 + 2 \text{ AlBr}_3$
 a. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 2.3 moles of $\text{Mg(NO}_3)_2$

$$2.3 \text{ mol Mg(NO}_3)_2 \times \frac{2 \text{ mol Al(NO}_3)_3}{3 \text{ mol Mg(NO}_3)_2} = \boxed{1.53 \text{ mol Al(NO}_3)_3}$$

- b. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 0.78 moles of AlBr_3

$$0.78 \text{ mol AlBr}_3 \times \frac{2 \text{ mol Al(NO}_3)_3}{2 \text{ mol AlBr}_3} = \boxed{0.78 \text{ mol Al(NO}_3)_3}$$

- c. Find the number of moles MgBr_2 that reacts with 1.05 moles of $\text{Al(NO}_3)_3$

$$1.05 \text{ mol Al(NO}_3)_3 \times \frac{3 \text{ mol MgBr}_2}{2 \text{ mol Al(NO}_3)_3} = \boxed{1.58 \text{ mol MgBr}_2}$$

- 4) In the reaction below, how many moles of carbon dioxide are produced if 19.65 moles of water are produced?
 $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$

$$19.65 \text{ moles H}_2\text{O} \times \frac{4 \text{ mol CO}_2}{6 \text{ mol H}_2\text{O}} = \boxed{13.1 \text{ mol CO}_2}$$

Stoichiometry Mole Practice Problems Answer Key

Heather Hattori, Richard H. Langley



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ChemDiscovery Teacher Edition Olga I. Agapova, 2002 Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material *The Practice of Chemistry* Donald J. Wink, Sharon

Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB **Introduction to**

General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE *MCAT General Chemistry Review 2018-2019* Kaplan Test Prep, 2017-07-04 Kaplan's MCAT Complete 7 Book Set Subject Review has all the information and strategies you need to score higher on the MCAT These books feature more practice than any other guide plus targeted strategy review opportunities for self analysis and thorough information on all of the critical thinking skills necessary for MCAT success from the creators of the 1 MCAT prep course From publisher's description

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Table of Contents Stoichiometry Mole Practice Problems Answer Key

1. Understanding the eBook Stoichiometry Mole Practice Problems Answer Key
 - The Rise of Digital Reading Stoichiometry Mole Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Mole Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Mole Practice Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Mole Practice Problems Answer Key
 - Personalized Recommendations
 - Stoichiometry Mole Practice Problems Answer Key User Reviews and Ratings
 - Stoichiometry Mole Practice Problems Answer Key and Bestseller Lists
5. Accessing Stoichiometry Mole Practice Problems Answer Key Free and Paid eBooks
 - Stoichiometry Mole Practice Problems Answer Key Public Domain eBooks
 - Stoichiometry Mole Practice Problems Answer Key eBook Subscription Services
 - Stoichiometry Mole Practice Problems Answer Key Budget-Friendly Options
6. Navigating Stoichiometry Mole Practice Problems Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Stoichiometry Mole Practice Problems Answer Key Compatibility with Devices
 - Stoichiometry Mole Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Mole Practice Problems Answer Key
 - Highlighting and Note-Taking Stoichiometry Mole Practice Problems Answer Key
 - Interactive Elements Stoichiometry Mole Practice Problems Answer Key
 8. Staying Engaged with Stoichiometry Mole Practice Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Mole Practice Problems Answer Key
 9. Balancing eBooks and Physical Books Stoichiometry Mole Practice Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Mole Practice Problems Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Mole Practice Problems Answer Key
 - Setting Reading Goals Stoichiometry Mole Practice Problems Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Mole Practice Problems Answer Key
 - Fact-Checking eBook Content of Stoichiometry Mole Practice Problems Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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