

Mole to Mole Stoichiometry Worksheet

- 1) In the reaction below, how many moles of aluminum will be used when reacted with 1.33 moles of oxygen?
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$

$$1.33 \text{ mol O}_2 \times \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = \boxed{1.77 \text{ mol Al}}$$

- 2) How many moles of hydrogen will be produced when reacted with 0.0204 moles of sodium in the reaction?
 $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + \text{H}_2$

$$0.0204 \text{ mol Na} \times \frac{1 \text{ mol H}_2}{2 \text{ mol Na}} = \boxed{0.0102 \text{ mol H}_2}$$

- 3) Using the equation $2 \text{ Al(NO}_3)_3 + 3 \text{ MgBr}_2 \rightarrow 3 \text{ Mg(NO}_3)_2 + 2 \text{ AlBr}_3$
 a. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 2.3 moles of $\text{Mg(NO}_3)_2$

$$2.3 \text{ mol Mg(NO}_3)_2 \times \frac{2 \text{ mol Al(NO}_3)_3}{3 \text{ mol Mg(NO}_3)_2} = \boxed{1.53 \text{ mol Al(NO}_3)_3}$$

- b. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 0.78 moles of AlBr_3

$$0.78 \text{ mol AlBr}_3 \times \frac{2 \text{ mol Al(NO}_3)_3}{2 \text{ mol AlBr}_3} = \boxed{0.78 \text{ mol Al(NO}_3)_3}$$

- c. Find the number of moles MgBr_2 that reacts with 1.05 moles of $\text{Al(NO}_3)_3$

$$1.05 \text{ mol Al(NO}_3)_3 \times \frac{3 \text{ mol MgBr}_2}{2 \text{ mol Al(NO}_3)_3} = \boxed{1.58 \text{ mol MgBr}_2}$$

- 4) In the reaction below, how many moles of carbon dioxide are produced if 19.65 moles of water are produced?
 $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$

$$19.65 \text{ moles H}_2\text{O} \times \frac{4 \text{ mol CO}_2}{6 \text{ mol H}_2\text{O}} = \boxed{13.1 \text{ mol CO}_2}$$

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Terry C. Jones



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