

Stoichiometry Limiting Reagent

Answers

- 1 Consider the following reaction:



- (a) How many grams of NH_3 can be produced from the reaction of 28 g of N_2 and 25 g of H_2 ?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34 \text{ g NH}_3$$

$$25 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 142 \text{ g NH}_3$$

34 g of NH_3 can be produced.

- (b) How much of the excess reagent is left over?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} \times \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} = 6 \text{ g H}_2 \text{ used}$$

$$25 \text{ g H}_2 - 6 \text{ g H}_2 = 19 \text{ g H}_2 \text{ remains}$$

- 2 Consider the following reaction:



What volume of H_2 at STP is produced from the reaction of 50 g of Mg and 75 g of HCl ?

$$50 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 46 \text{ L H}_2$$

$$75 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 23 \text{ L H}_2$$

23 L H_2 is produced.

- 3 Zinc and sulfur react to form zinc sulfide according to the following equation:



Suppose 25 g of zinc and 30 g of sulfur react. Which reagent is limiting?

$$25 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.4 \text{ g Zn}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol Zn}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 37.3 \text{ g ZnS}$$

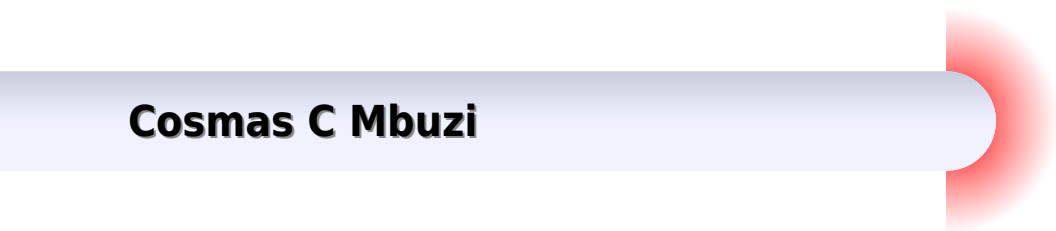
$$30 \text{ g S} \times \frac{1 \text{ mol S}}{32 \text{ g S}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol S}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 91.2 \text{ g ZnS}$$

Zn is the limiting reagent.

Stoichiometry Limiting Reagent Worksheet Answers

Chemistry If8766

Cosmas C Mbuzi



Stoichiometry Limiting Reagent Worksheet Answers Chemistry If8766:

Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities. Students in higher courses may find it a helpful revision and enhance their clarity. The book completely discusses the key topics in basic stoichiometry including the mole concept, reacting quantities, and empirical and molecular formulas. It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance. This is particularly useful to a beginner. Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems. The chapter also describes and applies the concept of a limiting reagent. Two methods of determining a limiting reagent are explained and illustrated. The concept of limiting reagent is extended to reacting volumes of gases with a short cut method. A short cut method for solving reacting quantities involving masses and volumes of gases is also given. Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas. A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner. In chapter 4 we discuss percentage yield and purity. A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances. A student will find these examples helpful in relating different formulas for percentage purity. The last chapter introduces a graphical method for reacting quantities. The method may provide a new way of looking at chemical reactions. Examples are given to illustrate the method including how it can be used to determine limiting reagents. It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course. Teachers may also find this book a good resource for their lessons in stoichiometry. *STOICHIOMETRY NARAYAN*

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