

Stoichiometry Limiting Reactant

Answers

- (1) How many grams of calcium phosphate can be produced from the reaction between 2.50 L of 0.250 M calcium chloride with an excess of phosphoric acid?



$$2.50 \text{ L CaCl}_2 \times \frac{0.250 \text{ mol CaCl}_2}{1 \text{ L CaCl}_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \times \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 64.6 \text{ g Ca}_3(\text{PO}_4)_2$$

- (2) How many milliliters of 1.50 M nitric acid is required to react with 100.0 g of cuprous oxide?



$$100 \text{ g Cu}_2\text{O} \times \frac{1 \text{ mol Cu}_2\text{O}}{143 \text{ g Cu}_2\text{O}} \times \frac{14 \text{ mol HNO}_3}{3 \text{ mol Cu}_2\text{O}} \times \frac{1000 \text{ ml HNO}_3}{1.5 \text{ mol HNO}_3} = 2180 \text{ ml HNO}_3$$

- (3) Consider the following unbalanced reaction:



Balance the reaction. 3.25 g of NH_3 is allowed to react with 3.5 g of O_2 .

- (a) Which reactant is the limiting reagent?

$$3.25 \text{ g NH}_3 \times \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 5.73 \text{ g NO}$$

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 2.63 \text{ g NO}$$

O_2 is the limiting reagent.

- (b) How many grams of NO are formed?

2.63 g of NO are formed.

- (c) How much excess reagent remains after the reaction?

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NH}_3}{5 \text{ mol O}_2} \times \frac{17.03 \text{ g NH}_3}{1 \text{ mol NH}_3} = 1.49 \text{ g NH}_3 \text{ used}$$

$$3.25 \text{ g NH}_3 - 1.49 \text{ g NH}_3 = 1.76 \text{ g NH}_3 \text{ remains.}$$

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Table of Contents Stoichiometry Limiting Reagent Answer Key

1. Understanding the eBook Stoichiometry Limiting Reagent Answer Key
 - The Rise of Digital Reading Stoichiometry Limiting Reagent Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Limiting Reagent Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Limiting Reagent Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Limiting Reagent Answer Key
 - Personalized Recommendations
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 - Stoichiometry Limiting Reagent Answer Key and Bestseller Lists
5. Accessing Stoichiometry Limiting Reagent Answer Key Free and Paid eBooks

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 - Stoichiometry Limiting Reagent Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Limiting Reagent Answer Key
 - Highlighting and Note-Taking Stoichiometry Limiting Reagent Answer Key
 - Interactive Elements Stoichiometry Limiting Reagent Answer Key
- 8. Staying Engaged with Stoichiometry Limiting Reagent Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Limiting Reagent Answer Key
- 9. Balancing eBooks and Physical Books Stoichiometry Limiting Reagent Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Limiting Reagent Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Limiting Reagent Answer Key
 - Setting Reading Goals Stoichiometry Limiting Reagent Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Limiting Reagent Answer Key
 - Fact-Checking eBook Content of Stoichiometry Limiting Reagent Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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