

Lab: Baking Soda Stoichiometry

Introduction: When baking soda (NaHCO_3) is an ingredient in your recipe, its purpose is to make the batter rise and produce a product with a light and fluffy texture. The decomposition reaction is shown below:



Our goal will be to predict how much sodium carbonate (Na_2CO_3) is produced when baking soda decomposes.

Materials: [Lab Video \(use to fill in data table\)](#)

Procedure:

1. Use the lab video to complete the data table below.
2. Calculate the mass of the NaHCO_3 used in the reaction and the experimental mass of Na_2CO_3 that was produced. **You need to show work in the "calculation" table for how you got your answer. Typing it out is okay for the calculation table section.**
3. Answer the 2 conclusion questions at the end.

Data table

Mass of crucible (g)	12.74 g
Mass of crucible and NaHCO_3 (g)	14.33 g
Mass of crucible and Na_2CO_3 (g)	13.75 g

Calculations Table (show work in each box)

Mass of NaHCO_3 (g)	1.59 g
Experimental (actual) mass of Na_2CO_3 (g)	1.01 g

Conclusion Questions: Show all work, handwritten, using the correct steps/methods shown during class. You will need to upload an image of your work.

1. Use your starting mass of NaHCO_3 and the balanced chemical equation above to calculate the mass of Na_2CO_3 that **should** be produced by the decomposition reaction. This is the theoretical yield.

Stoichiometry Lab Answer Key Baking Soda

James Joseph Gallagher



Stoichiometry Lab Answer Key Baking Soda:

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Table of Contents Stoichiometry Lab Answer Key Baking Soda

1. Understanding the eBook Stoichiometry Lab Answer Key Baking Soda
 - The Rise of Digital Reading Stoichiometry Lab Answer Key Baking Soda
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Lab Answer Key Baking Soda
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Lab Answer Key Baking Soda
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Lab Answer Key Baking Soda
 - Personalized Recommendations
 - Stoichiometry Lab Answer Key Baking Soda User Reviews and Ratings
 - Stoichiometry Lab Answer Key Baking Soda and Bestseller Lists
5. Accessing Stoichiometry Lab Answer Key Baking Soda Free and Paid eBooks
 - Stoichiometry Lab Answer Key Baking Soda Public Domain eBooks
 - Stoichiometry Lab Answer Key Baking Soda eBook Subscription Services
 - Stoichiometry Lab Answer Key Baking Soda Budget-Friendly Options
6. Navigating Stoichiometry Lab Answer Key Baking Soda eBook Formats

- ePub, PDF, MOBI, and More
 - Stoichiometry Lab Answer Key Baking Soda Compatibility with Devices
 - Stoichiometry Lab Answer Key Baking Soda Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Lab Answer Key Baking Soda
 - Highlighting and Note-Taking Stoichiometry Lab Answer Key Baking Soda
 - Interactive Elements Stoichiometry Lab Answer Key Baking Soda
 8. Staying Engaged with Stoichiometry Lab Answer Key Baking Soda
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Lab Answer Key Baking Soda
 9. Balancing eBooks and Physical Books Stoichiometry Lab Answer Key Baking Soda
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Lab Answer Key Baking Soda
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Lab Answer Key Baking Soda
 - Setting Reading Goals Stoichiometry Lab Answer Key Baking Soda
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Lab Answer Key Baking Soda
 - Fact-Checking eBook Content of Stoichiometry Lab Answer Key Baking Soda
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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