

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Stoichiometry Guided Study Work Answers

James E. Brady, Fred Senese



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988

Popular Science, 1877-06 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better

Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

Resources in Education, 1995

General Chemistry, Study Guide James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with

greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems *Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb* Wade A. Freeman, 1986 *Student Study Guide to accompany Chemistry* Martin Silberberg, Libby Weberg, 2005-01-06 *Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl* Paul B. Kelter, 1993 *Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book *The Extraordinary Chemistry of Ordinary Things, Study Guide* Carl H. Snyder, David F. Dever, 1995-01-19 Examines the chemistry of the substances of our everyday world Our daily lives are immersed in chemicals an effective way to teach and learn chemistry is by examining the goods and substances that we use in our daily lives and that affect us and our environment Progress in Ecological Stoichiometry Dedmer B. Van de Waal, James J. Elser, Adam C. Martiny, Robert W. Sterner, James B. Cotner, 2018 Ecological stoichiometry concerns the way that the elemental composition of organisms shapes their ecology It deals with the balance or imbalance of elemental ratios and how that affects organism growth nutrient cycling and the interactions with the biotic and abiotic worlds The elemental composition of organisms is a set of constraints through which all the Earth's biogeochemical cycles must pass All organisms consume nutrients and acquire compounds from the environment proportional to their needs Organismal elemental needs are determined in turn by the energy required to live and grow the physical and chemical constraints of their environment and their requirements for relatively large polymeric biomolecules such as RNA DNA lipids and proteins as well as for structural needs including stems bones shells etc These materials together constitute most of the biomass of living organisms Although there may be little variability in elemental ratios of many of these biomolecules changing the proportions of different biomolecules can have important effects on organismal elemental composition Consequently the variation in elemental composition both within and across organisms can be tremendous which has important implications for Earth's biogeochemical cycles It has been over a decade since the publication of Sterner and Elser's book *Ecological Stoichiometry* 2002 In the intervening years hundreds of papers on stoichiometric topics ranging from evolution and regulation of nutrient content in organisms to the role of stoichiometry in populations communities ecosystems and global biogeochemical dynamics have been published Here we present a collection of contributions from the broad scientific community to highlight recent insights in the field of Ecological Stoichiometry

Study Guide for Mortimer's Chemistry, Sixth Edition Donald W. Shive, Louise E. Shive, 1986

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1. Understanding the eBook Stoichiometry Guided Study Work Answers
 - The Rise of Digital Reading Stoichiometry Guided Study Work Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Guided Study Work Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Guided Study Work Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Guided Study Work Answers
 - Personalized Recommendations
 - Stoichiometry Guided Study Work Answers User Reviews and Ratings
 - Stoichiometry Guided Study Work Answers and Bestseller Lists
5. Accessing Stoichiometry Guided Study Work Answers Free and Paid eBooks
 - Stoichiometry Guided Study Work Answers Public Domain eBooks
 - Stoichiometry Guided Study Work Answers eBook Subscription Services
 - Stoichiometry Guided Study Work Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Stoichiometry Guided Study Work Answers Compatibility with Devices
 - Stoichiometry Guided Study Work Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Guided Study Work Answers
 - Highlighting and Note-Taking Stoichiometry Guided Study Work Answers
 - Interactive Elements Stoichiometry Guided Study Work Answers
8. Staying Engaged with Stoichiometry Guided Study Work Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Guided Study Work Answers
9. Balancing eBooks and Physical Books Stoichiometry Guided Study Work Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Guided Study Work Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Guided Study Work Answers
 - Setting Reading Goals Stoichiometry Guided Study Work Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Guided Study Work Answers
 - Fact-Checking eBook Content of Stoichiometry Guided Study Work Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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