

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before $\underline{2 \text{ mol}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-2 \text{ mol}}$ $\underline{+2 \text{ mol}}$

After $\underline{0 \text{ mol}}$ $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before $\underline{9.94 \text{ mol}}$ $\underline{x \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-9.94 \text{ mol}}$ $\underline{+9.94 \text{ mol}}$

After $\underline{0 \text{ mol}}$ $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to

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Table of Contents Stoichiometry 2 Percent Yield Answers

1. Understanding the eBook Stoichiometry 2 Percent Yield Answers
 - The Rise of Digital Reading Stoichiometry 2 Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry 2 Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry 2 Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry 2 Percent Yield Answers
 - Personalized Recommendations
 - Stoichiometry 2 Percent Yield Answers User Reviews and Ratings
 - Stoichiometry 2 Percent Yield Answers and Bestseller Lists
5. Accessing Stoichiometry 2 Percent Yield Answers Free and Paid eBooks
 - Stoichiometry 2 Percent Yield Answers Public Domain eBooks
 - Stoichiometry 2 Percent Yield Answers eBook Subscription Services
 - Stoichiometry 2 Percent Yield Answers Budget-Friendly Options
6. Navigating Stoichiometry 2 Percent Yield Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Stoichiometry 2 Percent Yield Answers Compatibility with Devices
 - Stoichiometry 2 Percent Yield Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry 2 Percent Yield Answers
 - Highlighting and Note-Taking Stoichiometry 2 Percent Yield Answers
 - Interactive Elements Stoichiometry 2 Percent Yield Answers
 8. Staying Engaged with Stoichiometry 2 Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry 2 Percent Yield Answers
 9. Balancing eBooks and Physical Books Stoichiometry 2 Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry 2 Percent Yield Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry 2 Percent Yield Answers
 - Setting Reading Goals Stoichiometry 2 Percent Yield Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry 2 Percent Yield Answers
 - Fact-Checking eBook Content of Stoichiometry 2 Percent Yield Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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