Stoichiometry

Solve the following problems:

- Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)
- a) Translate the problem to chemical equation.

$$4FeS + 7O_2 \longrightarrow 2Fe_2O_2 + 4SO_2$$

Reactant side	Product side
4 Fe atoms	4 Fe atoms
4 S atoms	4 S atoms
7(2) =14 O atoms	2(3) +4(2) =14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given: m FeS = 2.0 gReq'd. = $m O_2$

Solution: molar mass of FeS = 88 g/molmole ratio = 4 mol FeS: 7 mol O_2 molar mass of $O_2 = 32 \text{ g/mol}$

$$m_{O_2} = 2.0 g_{FeS} \ \left(\frac{1 \, \text{mot}_{Fes}}{88 g_{FeS}} \right) \left(\frac{7 \, \text{mot}_{O_2}}{4 \, \text{mot}_{FeS}} \right) \left(\frac{32 g_{O_2}}{1 \, \text{mot}_{O_2}} \right) = 1.3 g_{O_2}$$

- Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)
 - a) Write the balanced chemical equation.

 Reactant side
 Product side

 4 Fe atoms
 4 Fe atoms

 3(2) =6 O atoms
 2(3) =6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

Product to each reactant:

4 mol Fe: 3 mol O_2

2 mol FeO₂:4 mol Fe

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Table of Contents Stoichiometry Supplemental Problems Answer Key

- 1. Understanding the eBook Stoichiometry Supplemental Problems Answer Key
 - The Rise of Digital Reading Stoichiometry Supplemental Problems Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Stoichiometry Supplemental Problems Answer Key
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Supplemental Problems Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Stoichiometry Supplemental Problems Answer Key
 - Personalized Recommendations
 - Stoichiometry Supplemental Problems Answer Key User Reviews and Ratings
 - Stoichiometry Supplemental Problems Answer Key and Bestseller Lists

- 5. Accessing Stoichiometry Supplemental Problems Answer Key Free and Paid eBooks
 - Stoichiometry Supplemental Problems Answer Key Public Domain eBooks
 - Stoichiometry Supplemental Problems Answer Key eBook Subscription Services
 - Stoichiometry Supplemental Problems Answer Key Budget-Friendly Options
- 6. Navigating Stoichiometry Supplemental Problems Answer Key eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Stoichiometry Supplemental Problems Answer Key Compatibility with Devices
 - Stoichiometry Supplemental Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Supplemental Problems Answer Key
 - Highlighting and Note-Taking Stoichiometry Supplemental Problems Answer Key
 - o Interactive Elements Stoichiometry Supplemental Problems Answer Key
- 8. Staying Engaged with Stoichiometry Supplemental Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Supplemental Problems Answer Key
- 9. Balancing eBooks and Physical Books Stoichiometry Supplemental Problems Answer Key
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Stoichiometry Supplemental Problems Answer Key
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Supplemental Problems Answer Key
 - Setting Reading Goals Stoichiometry Supplemental Problems Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Supplemental Problems Answer Key
 - Fact-Checking eBook Content of Stoichiometry Supplemental Problems Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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