

# ANSWERS:

## Stoichiometry Quiz SCH3U Names

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### Assessment of Learning



1. Answer the following questions based on the balanced chemical equation above.

a) If 23.5 g of  $\text{C}_{10}\text{H}_{18}$  is mixed with  $3.76 \times 10^{23}$  molecules of oxygen, determine the limiting reactant. (show all work) (3T, 1C)

ANS: oxygen is limiting reactant

b) What mass of  $\text{H}_2\text{O}$  is expected in part a? (1T, 1C)

ANS: 6.99 g water

c) How many molecules  $\text{CO}_2$  of are expected in part a? (3T, 1C)

ANS:  $2.59 \times 10^{23}$  molecules of  $\text{CO}_2$

d) What mass of the excess reactant is left over after the reaction? (3T, 1C)

ANS: 17.5 g  $\text{C}_{10}\text{H}_{18}$

2. During a lab a student reacts 18.4 g of nitrogen with excess oxygen.



a. Balance the chemical equation above. (2A)

ANS:  $2\text{N}_2 + 5\text{O}_2 \rightarrow 2\text{N}_2\text{O}_5$

b. Calculate the molecules of dinitrogen pentoxide that form, if the reaction proceeds with a 58.2% yield (4A)

ANS:  $2.30 \times 10^{23}$  molecules of  $\text{N}_2\text{O}_5$

# Stoichiometry Quiz For Chemistry With Answers

**Yijin Wang**



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