

STOICHIOMETRY PROBLEMS

1. The combustion of a sample of butane, C_4H_{10} (lighter fluid), produced 2.46 grams of water.



(a) How many moles of water formed?

$$2.46 \text{ g H}_2\text{O} \times (1 \text{ mol H}_2\text{O}/18 \text{ g H}_2\text{O}) = 0.137 \text{ mol H}_2\text{O}$$

(b) How many moles of butane burned?

$$0.137 \text{ mol H}_2\text{O} \times (2 \text{ mol C}_4\text{H}_{10}/10 \text{ mol H}_2\text{O}) = 0.027 \text{ mol C}_4\text{H}_{10}$$

(c) How many grams of butane burned?

$$0.027 \text{ mol C}_4\text{H}_{10} \times (58.12 \text{ g C}_4\text{H}_{10}/1 \text{ mol C}_4\text{H}_{10}) = 7.94 \text{ g C}_4\text{H}_{10}$$

(d) How much oxygen was used up in moles?

$$0.137 \text{ mol H}_2\text{O} \times (13 \text{ mol O}_2/10 \text{ mol H}_2\text{O}) = 0.1781 \text{ mol O}_2$$

(e) How much oxygen was used up in grams?

$$0.1781 \text{ mol O}_2 \times (16 \text{ g O}_2/1 \text{ mol O}_2) = 2.85 \text{ g O}_2$$

2. Sodium metal burns in air according to the balanced reaction below.



(a) How many moles of oxygen are needed to react with 9.5 g of sodium completely?

$$9.5 \text{ g Na} \times (1 \text{ mol Na}/23 \text{ g Na}) \times (1 \text{ mol O}_2/4 \text{ mol Na}) = 0.103 \text{ mol O}_2$$

(b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?

$$12.5 \text{ g Na}_2\text{O} \times (1 \text{ mol Na}_2\text{O}/62 \text{ g Na}_2\text{O}) \times (4 \text{ mol Na}/2 \text{ mol Na}_2\text{O}) \times (23 \text{ g Na}/1 \text{ mol Na}) = 9.27 \text{ g Na}$$

3. In the combustion of 54.5 g of butane (C_4H_{10}), how many grams of CO_2 are produced? Write and balance the equation before solving.



$$54.5 \text{ g C}_4\text{H}_{10} \times (1 \text{ mol C}_4\text{H}_{10}/58 \text{ g C}_4\text{H}_{10}) \times 8 \text{ mol CO}_2/2 \text{ mol C}_4\text{H}_{10} \times (44 \text{ g CO}_2/1 \text{ mol CO}_2) = 178 \text{ g CO}_2$$

Stoichiometry Quiz Answer Key

Liyong Dong



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