

Name: Key

Class: _____

Date: _____

Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$, how many grams of magnesium ^B nitrate are produced from 8.00 mol of nitric acid (HNO_3)? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$, how many moles of potassium ^B chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of water are produced from 5.22 liters of oxygen gas? ^B A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of hydrogen ^B are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$, how many moles of carbonic ^B acid can be produced from 97.2 g of CO_2 ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1
Feeling VERY confused

2
Not sure that I get this

3
I think I am catching on

4
Doing great! Bring it on!

Stoichiometry Problems 1 Answers

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Table of Contents Stoichiometry Problems 1 Answers

1. Understanding the eBook Stoichiometry Problems 1 Answers
 - The Rise of Digital Reading Stoichiometry Problems 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Problems 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Problems 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Problems 1 Answers
 - Personalized Recommendations
 - Stoichiometry Problems 1 Answers User Reviews and Ratings
 - Stoichiometry Problems 1 Answers and Bestseller Lists

5. Accessing Stoichiometry Problems 1 Answers Free and Paid eBooks
 - Stoichiometry Problems 1 Answers Public Domain eBooks
 - Stoichiometry Problems 1 Answers eBook Subscription Services
 - Stoichiometry Problems 1 Answers Budget-Friendly Options
6. Navigating Stoichiometry Problems 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Problems 1 Answers Compatibility with Devices
 - Stoichiometry Problems 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Problems 1 Answers
 - Highlighting and Note-Taking Stoichiometry Problems 1 Answers
 - Interactive Elements Stoichiometry Problems 1 Answers
8. Staying Engaged with Stoichiometry Problems 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Problems 1 Answers
9. Balancing eBooks and Physical Books Stoichiometry Problems 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Problems 1 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Problems 1 Answers
 - Setting Reading Goals Stoichiometry Problems 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Problems 1 Answers
 - Fact-Checking eBook Content of Stoichiometry Problems 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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