

Stoichiometry: Practice Assessment

Name: _____

5888 - Using a balanced equation, given the grams of reactant, calc. the grams of product



Balance the equation provided. Calculate the number of moles of sodium bromide that would be produced if 25.5 moles of gallium (III) bromide are used (with an excess of sodium sulfite)

$$25.5\text{mol GaBr}_3 \times 6\text{mol NaBr}/2\text{mol GaBr}_3 = 75.75\text{mol NaBr}$$

2) Tin (II) oxide reacts with nitrogen trifluoride to produce tin (II) fluoride and dinitrogen trioxide. Using the following reaction, calculate the grams produced of each product from 50.0 grams of nitrogen trifluoride.



$$50\text{g NF}_3 \times 1\text{mol NF}_3 / 71.001\text{g NF}_3 = 0.704\text{mol NF}_3 \times 2\text{mol SnF}_2 / 2\text{mol NF}_3 \times 156.69\text{g SnF}_2 / 1\text{mol SnF}_2 = 165.5\text{g SnF}_2$$

SnF₂

$$= 0.704\text{mol NF}_3 \times 1\text{mol N}_2\text{O}_3 / 2\text{mol NF}_3 \times 76.014\text{g N}_2\text{O}_3 / 1\text{mol N}_2\text{O}_3 =$$

26.75g N₂O₃

3) A piece of copper wire is placed in silver nitrate. Silver and aqueous copper (II) nitrate is produced. The initial mass of the copper before the reaction was 2.74g and the mass remaining after the reaction was 2.32g. What is the mass of silver produced?



$$2.74\text{g} - 2.32\text{g} = 0.42\text{g}$$

$$0.42\text{g Cu} \times 1\text{mol Cu} / 63.55\text{g Cu} \times 2\text{mol Ag} / 1\text{mol Cu} \times 107.87\text{g Ag} / 1\text{mol Ag} = 1.43\text{g Ag}$$

5889 - After calculating theoretical yield in grams, calculate the percent yield. Given experimental data, determine the limiting reactant and the excess reactant

a) Through experimentation, you observe the reaction of iron (III) phosphate with sodium sulfate to make iron (III) sulfate and sodium phosphate. You began the experiment with 25g of iron (III) phosphate and 182g of sodium sulfate. How many grams of iron (III) sulfate should you have been able to produce? What is the percent yield if 18.5g of iron (III) sulfate are actually made?



$$25\text{g FePO}_4 \times 1\text{mol FePO}_4 / 159.815\text{g FePO}_4 = 0.156\text{mol FePO}_4 / 2\text{mol per rxn} = 0.083\text{rxns (LR)}$$

$$182\text{g Na}_2\text{SO}_4 \times 1\text{mol Na}_2\text{SO}_4 / 142.04\text{g Na}_2\text{SO}_4 = 1.28\text{mol Na}_2\text{SO}_4 / 3\text{mol per rxn} = 0.427\text{rxns}$$

$$0.083\text{mol FePO}_4 \times 1\text{mol Fe}_2(\text{SO}_4)_3 / 2\text{mol FePO}_4 \times 399.873\text{g Fe}_2(\text{SO}_4)_3 / 1\text{mol Fe}_2(\text{SO}_4)_3 = 33.19\text{g}$$

$$(18.5\text{g} / 33.19\text{g}) \times 100 = 56\%$$

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Ralph Thomas Myers



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Holt Chemistry Ralph Thomas Myers, 2004 **Books in Print Supplement** , 2002

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Table of Contents Stoichiometry Packet 9 Answer Key Guided Reading

1. Understanding the eBook Stoichiometry Packet 9 Answer Key Guided Reading
 - The Rise of Digital Reading Stoichiometry Packet 9 Answer Key Guided Reading
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Packet 9 Answer Key Guided Reading
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Packet 9 Answer Key Guided Reading
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Packet 9 Answer Key Guided Reading

- Personalized Recommendations
 - Stoichiometry Packet 9 Answer Key Guided Reading User Reviews and Ratings
 - Stoichiometry Packet 9 Answer Key Guided Reading and Bestseller Lists
5. Accessing Stoichiometry Packet 9 Answer Key Guided Reading Free and Paid eBooks
 - Stoichiometry Packet 9 Answer Key Guided Reading Public Domain eBooks
 - Stoichiometry Packet 9 Answer Key Guided Reading eBook Subscription Services
 - Stoichiometry Packet 9 Answer Key Guided Reading Budget-Friendly Options
 6. Navigating Stoichiometry Packet 9 Answer Key Guided Reading eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Packet 9 Answer Key Guided Reading Compatibility with Devices
 - Stoichiometry Packet 9 Answer Key Guided Reading Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Packet 9 Answer Key Guided Reading
 - Highlighting and Note-Taking Stoichiometry Packet 9 Answer Key Guided Reading
 - Interactive Elements Stoichiometry Packet 9 Answer Key Guided Reading
 8. Staying Engaged with Stoichiometry Packet 9 Answer Key Guided Reading
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Packet 9 Answer Key Guided Reading
 9. Balancing eBooks and Physical Books Stoichiometry Packet 9 Answer Key Guided Reading
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Packet 9 Answer Key Guided Reading
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Packet 9 Answer Key Guided Reading
 - Setting Reading Goals Stoichiometry Packet 9 Answer Key Guided Reading
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Packet 9 Answer Key Guided Reading

- Fact-Checking eBook Content of Stoichiometry Packet 9 Answer Key Guided Reading
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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