



Name: Chelsey Williams Date: 3/31/22

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? 15 sticks of butter

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 1 CO 3 Fe 2 CO_2 3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe_2O_3)?

Problems solved: 0



Stoichiometry Lab Chemfiesta Answers

Jin-Ying Zhang



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A Stoichiometry Unit David Callaghan, 2004 *Stoichiometry Unit Project* Luann Marie Decker, 1998 *Stoichiometry* Sydney Young, 1908 **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction *You Can Do Chemistry* Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 **Chemistry Study Guide** Gregory Gebhart, 2017-09-21 Solutions to stoichiometry and conservation of mass problems Laboratory exercises

Stoichiometry B. I. Bhatt, Shuchen B. Thakore, 2010 *Stoichiometry* B I. Bhatt, 2010 **Improving Student Comprehension of Stoichiometric Concepts** Connie Lynn Bannick Kemner, 2007 **Stoichiometry, 4E** B. I. Bhatt, 2004-03 *Formulation and Stoichiometry* Emil J Margolis, **Bite-Sized Chemistry Calculations** John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry

and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

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