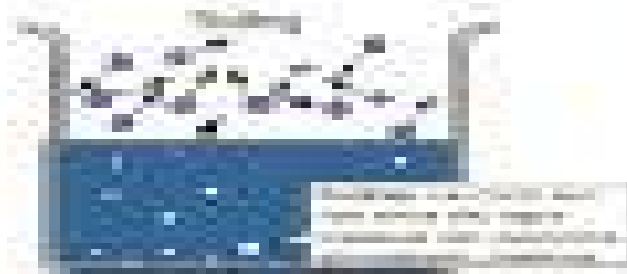
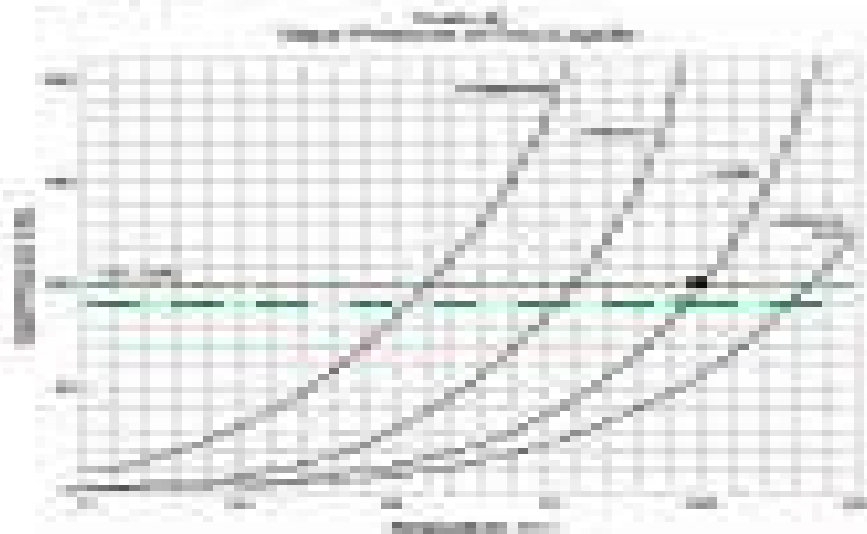


Liquid Pressure (Clarified)

Figure 1

The liquid pressure at a point is dependent on the depth down. At any particular depth, pressure against the walls of the container is the same (the fluid is isotropic) regardless of the fact at that "depth" they could be at the bottom of a container, subjected to some high atmospheric pressure that we can feel, or even inside some fluid at almost zero pressure. Consider a container, surface has solid and sides, depending on the structure of the container pressure:

Figure 1:



<https://www.youtube.com/watch?v=1234567890>

Liquid with free surface

liquid pressure is atmospheric pressure

The liquid pressure at a point is dependent on the depth down. At any particular depth, pressure against the walls of the container is the same (the fluid is isotropic) regardless of the fact at that "depth" they could be at the bottom of a container, subjected to some high atmospheric pressure that we can feel, or even inside some fluid at almost zero pressure. Consider a container, surface has solid and sides, depending on the structure of the container pressure:

Each liquid with free surface point where the liquid pressure is zero is the atmospheric pressure. Consider atmospheric pressure and depth of the liquid. The liquid pressure at a point is dependent on the depth down. At any particular depth, pressure against the walls of the container is the same (the fluid is isotropic) regardless of the fact at that "depth" they could be at the bottom of a container, subjected to some high atmospheric pressure that we can feel, or even inside some fluid at almost zero pressure. Consider a container, surface has solid and sides, depending on the structure of the container pressure:

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Francis Craig Krauskopf



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Vapor Pressures of Various Compounds at Low Temperatures G. A. Burrell, I. W. Robertson, 1916 *Science Citation Index*, 1992 Vols for 1964 have guides and journal lists Vapor Pressure and Vaporization of Petroleum Fractions Donald La Verne Katz, George Granger Brown, 1933 *Vapor Pressure Curves* W. N. Holmes, 1913 **The Vapor Pressures of Aqueous Solutions of Electrolytes** Edward Otto Heuse, 1915 **THE VAPOR PRESSURES OF 30 INORGANIC LIQUIDS BETWEEN ONE ATMOSPHERE AND THE CRITICAL POINT.**, 1963 Literature values of the vapor pressures of H₂ HD D₂ N₂ O₂ Ne Ar Kr Xe CO NO CH₄ N₂O CO₂ HCl HBr HI SF₆ H₂S Cl₂ Br₂ NH₃ C₂H₂ SO₂ SO₃ COCl₂ HF HCN N₂O₄ CCl₄ and SnCl₄ were compared with the Kirchhoff formula $\log P_K = A + \frac{B}{T - B}$ between the normal boiling point and the critical point where A and B are determined for each liquid Plots are presented of the deviations computed by the equation $\log P/P_K = b$ Based on accurately determined curves for 17 of these substances empirical hypotheses were developed to allow the prediction of the most probable course of the deviation curve in cases where the data appeared to be unreliable In addition where precise data were available the possibility of the existence of fine structure curvature was demonstrated The deviation curves were used to obtain tables of interpolated vapor pressure values which are in most cases more accurate by one order of magnitude than any now in the literature particularly where smoothed regression curves do not discriminate between random experimental errors and systematic deviations of an arbitrarily chosen formula from the real vapor pressure curve Since the deviation curve is invariably shown to approach the zero ordinate at the critical point with a large negative slope the correspondence between the critical pressure and temperature is fixed with a high degree of confidence A particular result of this work is that the critical pressure of NH₃ was shown to be about one atmosphere higher than the presently accepted value The new values for the critical points of CH₄ C₂N₂ HCN and Br₂ also resolved differences existing between various authorities in the literature auth *Vapor Pressure Curves for C, Ni, Fe, Cr, Mo, W, Cu, Ta, and Al Over the Pressure Range from 10⁻¹⁰ ATM to 1 ATM* A. R. Smider, 1952 **Vapor Pressure Curves for Systems Containing Alcohol, Ether and Water ...** Earl Albert Louder, Thomas Roland Brigg, Arthur Wesley Browne, 1924 **Vapor Pressure Data Analysis Methodology, Statistics, and Applications** Elwin C. Penski, 1992 A method of fitting vapor pressure temperature data to the Antoine equation is described along with a detailed statistical analysis of the fits and calculations of vapor pressures enthalpies of vaporization volatilities concentration of saturated vapors and boiling points with the Antoine constants It is shown that the number of digits required to accurately perform vapor pressure calculations is greater than previously believed The system permits calculation of the limits of error for any level of confidence A program is used to analyze and plot the available vapor pressure data for diethyl malonate Vapor pressure Antoine equation Volatility Temperature Statistics Plotting Enthalpy of Vaporization Concentration of saturated vapor Experimental methods Diethyl malonate Confidence limits Confidence level Vapor Pressures of Various Compounds at Low Temperatures Francis

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