

CHEMICAL EQUILIBRIUM

✓ $K_c = \frac{[B]}{[A]} = \text{mol L}^{-1}$

✓ $K_p = \frac{P_p}{P_r} = \text{Partial Pressure}$

✓ $K_x = \frac{X_p}{X_r} = \text{mole fraction}$

✓ $K_p = K_c(RT)^{\Delta n_g}$ where $\Delta n_g = n_p - n_r$

✓ Predicting the extent of reaction :

- $K_c > 10^3$ [Forward reaction is favoured.]
- $K_c < 10^{-3}$ [Reverse reaction is favoured.]
- $10^{-3} < K_c < 10^3$ [Both reactants and products are present in equilibrium]

✓ Free Energy Change (ΔG)

- If $\Delta G = 0$ then reversible reaction would be in equilibrium, $K_c = 1$
- If $\Delta G = (+)$ ve then equilibrium will be displaced in backward direction; $K_c < 1$
- If $\Delta G = (-)$ ve then equilibrium will shift in forward direction; $K_c > 1$

✓ (a) K_c unit $\rightarrow (\text{moles/lit})^{\Delta n_g}$

➤ ΔG of a reaction under any set of conditions is related to its value under standard conditions, i.e. ΔG° by the equation

$$\Delta G = \Delta G^\circ + 2.303 RT \log Q$$

➤ Under equilibrium conditions, for same number of moles of reactants and products

$$Q = K_p = K_c = K \text{ and } \Delta G = 0$$

$$\therefore \Delta G^\circ = -2.303 RT \log K$$

➤ We have replaced K_p by K called thermodynamic equilibrium constant

Significance of ΔG°

Significance of ΔG° can be explained from the following points

(i) If $\Delta G^\circ < 0$, $\log K > 0$, $\Rightarrow K > 1$

Hence, reaction is spontaneous in forward direction.

(ii) Hence $\Delta G^\circ < 0$, $\log K > 0$, $\Rightarrow K < 1$

Hence, reaction is non-spontaneous or a reaction proceeds in the forward direction to such a small extent that a very small amount of the product is formed.

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Camilla Rothe



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Chemistry McGraw-Hill Staff, 2001-03 **Lamefoot Mine Site, Supplement to the Kettle River Key Project Expansion**, 1994 **Science Books**, 1967 *Supplementary List of Publications of the National Bureau of Standards*, 1960 **Physical Chemistry** Keith James Laidler, John H. Meiser, 1982 **Scientific American Supplement**, 1917 **NASA Technical Paper**, 1984 **Salters Advanced Chemistry** George Burton, 2000-07-31 The texts in the Salters Advanced Chemistry series have been updated to match the specifications for A Level Chemistry from September 2000 This supplement pack is designed to help teachers to use the original editions of the texts until they can be replaced

Resources in Education, 1994 *Current Catalog* National Library of Medicine (U.S.), 1974 First multi year cumulation covers six years 1965-70 **Supplementary Readings for Chemical Bond Approach** Robert K. Fitzgerald, William Franklin Kieffer, 1960 Government-wide Index to Federal Research & Development Reports, 1967 Essentials of Chemical Reaction Engineering H. Scott Fogler, 2011 Accompanying DVD ROM contains many realistic interactive simulations **Phase Equilibria in Chemical Engineering** Stanley M. Walas, 2013-10-22 Phase Equilibria in Chemical Engineering is devoted to the thermodynamic basis and practical aspects of the calculation of equilibrium conditions of multiple phases that are pertinent to chemical engineering processes Efforts have been made throughout the book to provide guidance to adequate theory and practice The book begins with a long chapter on equations of state since it is intimately bound up with the development of thermodynamics Following material on basic thermodynamics and nonidealities in terms of fugacities and activities individual chapters are devoted to equilibria primarily between pairs of phases A few topics that do not fit into these categories and for which the state of the art is not yet developed quantitatively have been relegated to a separate chapter The chapter on chemical equilibria is pertinent since many processes involve simultaneous chemical and phase equilibria Also included are chapters on the evaluation of enthalpy and entropy changes of nonideal substances and mixtures and on experimental methods This book is intended as a reference and self study as well as a textbook either for full courses in phase equilibria or as a supplement to related courses in the chemical engineering curriculum Practicing engineers concerned with separation technology and process design also may find the book useful **Applied Mechanics Reviews**, 1949 Monthly Catalog of United States Government Publications, Cumulative Index United States. Superintendent of Documents, 1978 *Aeronautical Engineering: A Cumulative Index to a Continuing Bibliography (supplement 248)*, 1990 **Elements of Mechanics** Henry Anton Erikson, 1932 **Key to algebraical factors [&c.]**. Dorabji H. Vachha, 1898 **Water Quality: Management and Pollution Control Problems (Jerusalem Workshop Papers)** Samuel Harry Jenkins, 1973

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Table of Contents Supplemental Problems Answer Sheet Chemical Equilibrium

1. Understanding the eBook Supplemental Problems Answer Sheet Chemical Equilibrium
 - The Rise of Digital Reading Supplemental Problems Answer Sheet Chemical Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Supplemental Problems Answer Sheet Chemical Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Supplemental Problems Answer Sheet Chemical Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Supplemental Problems Answer Sheet Chemical Equilibrium
 - Personalized Recommendations
 - Supplemental Problems Answer Sheet Chemical Equilibrium User Reviews and Ratings

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- Supplemental Problems Answer Sheet Chemical Equilibrium and Bestseller Lists
- 5. Accessing Supplemental Problems Answer Sheet Chemical Equilibrium Free and Paid eBooks
 - Supplemental Problems Answer Sheet Chemical Equilibrium Public Domain eBooks
 - Supplemental Problems Answer Sheet Chemical Equilibrium eBook Subscription Services
 - Supplemental Problems Answer Sheet Chemical Equilibrium Budget-Friendly Options
- 6. Navigating Supplemental Problems Answer Sheet Chemical Equilibrium eBook Formats
 - ePub, PDF, MOBI, and More
 - Supplemental Problems Answer Sheet Chemical Equilibrium Compatibility with Devices
 - Supplemental Problems Answer Sheet Chemical Equilibrium Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Supplemental Problems Answer Sheet Chemical Equilibrium
 - Highlighting and Note-Taking Supplemental Problems Answer Sheet Chemical Equilibrium
 - Interactive Elements Supplemental Problems Answer Sheet Chemical Equilibrium
- 8. Staying Engaged with Supplemental Problems Answer Sheet Chemical Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Supplemental Problems Answer Sheet Chemical Equilibrium
- 9. Balancing eBooks and Physical Books Supplemental Problems Answer Sheet Chemical Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Supplemental Problems Answer Sheet Chemical Equilibrium
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Supplemental Problems Answer Sheet Chemical Equilibrium
 - Setting Reading Goals Supplemental Problems Answer Sheet Chemical Equilibrium
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Supplemental Problems Answer Sheet Chemical Equilibrium
 - Fact-Checking eBook Content of Supplemental Problems Answer Sheet Chemical Equilibrium
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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