# Table F Solubility Guidelines for Aqueous solutions

Ions That Form Soluble Compounds	Exceptions	Ions Tl Insoluble
Group 1 ions (Li+, Na+, etc.)	**	carbonate i
ammonium (NH <sub>4</sub> *)		chromate (
nitrate (NO <sub>3</sub> <sup>-</sup> )		
acetate (C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> or CH <sub>3</sub> COO <sup>-</sup> )		phosphate
hydrogen carbonate (HCO <sub>3</sub> <sup>-</sup> )		sulfide (S <sup>2</sup> -
chlorate (ClO <sub>3</sub> <sup>-</sup> )		hydroxide (
halides (Cl <sup>+</sup> , Br <sup>+</sup> , I <sup>+</sup> )	when combined with Ag*, Pb <sup>2</sup> *, or Hg <sub>2</sub> <sup>2</sup> *	
sulfates (SO <sub>4</sub> <sup>2-</sup> )	when combined with Ag*, Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> , or Pb <sup>2+</sup>	*compounds

Ions That Form Insoluble Compounds*	Exceptions	
carbonate (CO <sub>3</sub> <sup>2-</sup> )	when combined with Group I fons or ammonium $(NH_4^+)$	
chromate (CrO <sub>4</sub> <sup>2-</sup> )	when combined with Group I ions, $Ca^{2+}$ , $Mg^{2+}$ , or ammonium $(NH_4^+)$	
phosphate (PO <sub>4</sub> <sup>3-</sup> )	when combined with Group I tons or ammonium $(NH_4^+)$	
sulfide (S <sup>2-</sup> )	when combined with Group 1 ions or ammonium (NH <sub>4</sub> *)	
hydroxide (OH=)	when combined with Group 1 ions, $\mathrm{Ca^{2+}}$ , $\mathrm{Ba^{2+}}$ , $\mathrm{Sr^{2+}}$ , or ammonium $(\mathrm{NH_4^+})$	

<sup>\*</sup>compounds having very low solubility in H<sub>2</sub>O

# **Solubility Guidelines For Aqueqeus Solutions Worjsheet Answers**

**Atherton Seidell** 

# **Solubility Guidelines For Aquequus Solutions Worjsheet Answers:**

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