

Solubility Product, K_{sp} - Worksheet 1

1) Write the dissociation equation and the solubility product (K_{sp}) expression for the dissociation of each of the following substances.

- | | | |
|---------------------------------|-----------------------------------|-------------------------------------|
| a) PbSO_4 (s) | c) $\text{Pb}(\text{NO}_3)_2$ (s) | e) Na_2S (s) |
| b) $\text{Mg}(\text{OH})_2$ (s) | d) AlCl_3 (s) | f) $\text{Ca}_3(\text{PO}_4)_2$ (s) |

2) Write the solubility product expression for the dissociation of each of the following solids in water.

- | | | |
|----------------------|-----------------------|----------------------|
| a) sodium chloride | c) silver sulfide | e) magnesium sulfite |
| b) calcium carbonate | d) aluminum hydroxide | f) barium phosphate |

3) The solubility of NiS is 1.0×10^{-12} mol/L. Calculate the value of K_{sp} for this compound.

4) The solubility of calcium hydroxide ($\text{Ca}(\text{OH})_2$) is 0.02 mol/L. Calculate the solubility product constant for this substance.

5) The solubility of silver sulfide (Ag_2S) is 1.3×10^{-6} mol/L. Calculate the K_{sp} value.

6) Experiments show that 0.0059 g of SrCO_3 will dissolve in 1.0 L of water. What is the K_{sp} for strontium carbonate?

7) The solubility of cobalt (II) sulfide is 3.8×10^{-3} g/L. Calculate the value of K_{sp} .

Solubility Product, K_{sp} - Worksheet 2

1. Calculate the solubility (in mol/L) for each of the following compounds:

- | | |
|-----------------------------|--------------------------------|
| a. FeS | $K_{sp} = 3.7 \times 10^{-19}$ |
| b. $\text{Mg}(\text{OH})_2$ | $K_{sp} = 9.0 \times 10^{-12}$ |
| c. Ag_2S | $K_{sp} = 1.6 \times 10^{-49}$ |
| d. AgCN | $K_{sp} = 2.0 \times 10^{-12}$ |

2. What is the solubility of silver iodide in grams per litre? ($K_{sp} = 8.3 \times 10^{-17}$)

3. The K_{sp} value for BaSO_4 is 1.5×10^{-9} . Calculate the number of moles of this salt that will dissolve in 1000 mL of water.

4. Calculate the mass of CuS that will dissolve in 1.0 litre of water. $K_{sp} = 4.0 \times 10^{-38}$

5. Calculate the mass of barium carbonate that will dissolve in 100 mL of water. The K_{sp} of BaCO_3 is 8.1×10^{-9} .

Solubility Product Answers Chemfiesta

Sherry Lewin



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1. Understanding the eBook Solubility Product Answers Chemfiesta
 - The Rise of Digital Reading Solubility Product Answers Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Product Answers Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Product Answers Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Product Answers Chemfiesta
 - Personalized Recommendations
 - Solubility Product Answers Chemfiesta User Reviews and Ratings
 - Solubility Product Answers Chemfiesta and Bestseller Lists
5. Accessing Solubility Product Answers Chemfiesta Free and Paid eBooks
 - Solubility Product Answers Chemfiesta Public Domain eBooks
 - Solubility Product Answers Chemfiesta eBook Subscription Services
 - Solubility Product Answers Chemfiesta Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Solubility Product Answers Chemfiesta Compatibility with Devices
 - Solubility Product Answers Chemfiesta Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Solubility Product Answers Chemfiesta
 - Highlighting and Note-Taking Solubility Product Answers Chemfiesta
 - Interactive Elements Solubility Product Answers Chemfiesta
8. Staying Engaged with Solubility Product Answers Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Product Answers Chemfiesta
9. Balancing eBooks and Physical Books Solubility Product Answers Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Product Answers Chemfiesta
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solubility Product Answers Chemfiesta
 - Setting Reading Goals Solubility Product Answers Chemfiesta
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solubility Product Answers Chemfiesta
 - Fact-Checking eBook Content of Solubility Product Answers Chemfiesta
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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