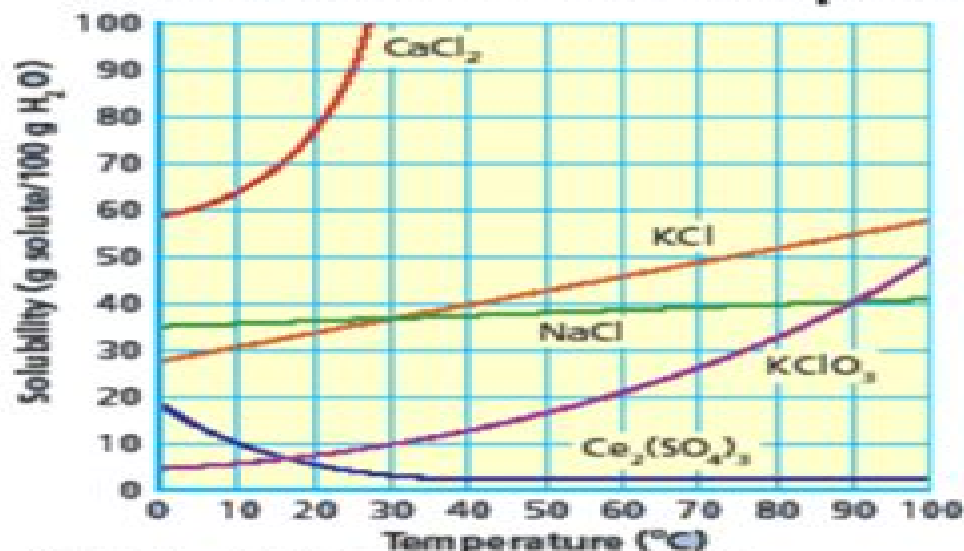


Directions: Use the graph below to answer the following questions. If the question requires a calculation, **SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER**. No work, no credit, no kidding! Remember units.

Solubilities as a Function of Temperature



Types of Solutions

- Points that collectively make up the solubility curve (points **ON** the curve) represent _____ solutions.
- Points that are **BELOW** the curve represent _____ solutions.
- Points **ABOVE** the solubility curve represent _____ solutions, and the difference between the point above the curve and on the curve represents the amount of solute which will precipitate out.

1. a. What is the solubility of calcium chloride at 5°C? _____ g/100g of H₂O
 b. What is the solubility of calcium chloride at 25°C? _____ g/100g of H₂O
 c. What is happening to the solubility and temperature?
 Solubility is _____ Increasing or Decreasing _____ as temperature is _____ Increasing or Decreasing _____.
 What is this type of relationship called? _____ direct or indirect _____.
 This occurs for which type of solutes? _____ gasses or solids _____.
2. a. What is the solubility of cerium sulfate at 10°C? _____ g/100g of H₂O
 b. What is the solubility of cerium sulfate at 50°C? _____ g/100g of H₂O
 c. What is happening to the solubility and temperature?
 Solubility is _____ increasing or decreasing _____ as temperature is _____ increasing or decreasing _____.
 What is this type of relationship called? _____ direct or indirect _____.
 Is this solution saturated, unsaturated, or supersaturated?
3. a. At 90°C, 10 g of potassium chlorate is dissolved in 100. g of water.
 Is this solution saturated, unsaturated, or supersaturated?
 b. How do you know?

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