

CHEM 120 Exam 2: Chapters 5,6,7,8 (With Answer Key) - Chamberlain College of Nursing

Examination Study Guide 2023

Exam 2

Total Points: 60

PART 1: Multiple Choice questions: Q 1-24. (2 point each)

- Chemical reactions (equations) show the following
 - Reactants
 - Products
 - Physical state of each (gas or liquid or solid)
 - All of the above
- The following is true for a mole
 - It is also known as Avagadro's number
 - It contains 6.022×10^{23} molecules
 - It contains mass equivalent to molecular weight
 - All of the above are true
- In balancing chemical equations, the coefficients are used to
 - Show number of moles
 - Show volume of each in liters
 - Show amount of each in grams
 - Show temperature at which the substance is stable
- $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$ means that
 - 2 grams of H_2 g and 1 gram of O_2 gas combine to form 2 grams of water.
 - 2 moles of H_2 g and 1 mole of O_2 gas combine to form 2 moles of water.
 - 2 grams of H_2 g and 1 gram of O_2 gas combine to form 2 moles of water.
 - 2 molecules of H_2 g and 1 molecule of O_2 gas combine to form 2 molecules of water.
- $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$. In this equation, correct method to calculate molecular mass of water is
 - $2[2(\text{mass of hydrogen}) + \text{mass of oxygen}]$
 - $2(\text{mass of hydrogen gas}) + \text{mass of oxygen gas}$
 - $2(\text{mass of hydrogen}) + \text{mass of oxygen}$

Todins Unit Exam Chem Answer Key

**Gretchen Bitterlin,Dennis
Johnson,Donna Price,Sylvia Ramirez**

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Table of Contents Todins Unit Exam Chem Answer Keg

1. Understanding the eBook Todins Unit Exam Chem Answer Keg
 - The Rise of Digital Reading Todins Unit Exam Chem Answer Keg
 - Advantages of eBooks Over Traditional Books
2. Identifying Todins Unit Exam Chem Answer Keg
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Todins Unit Exam Chem Answer Keg
 - User-Friendly Interface
4. Exploring eBook Recommendations from Todins Unit Exam Chem Answer Keg
 - Personalized Recommendations
 - Todins Unit Exam Chem Answer Keg User Reviews and Ratings
 - Todins Unit Exam Chem Answer Keg and Bestseller Lists

5. Accessing Todins Unit Exam Chem Answer Keg Free and Paid eBooks
 - Todins Unit Exam Chem Answer Keg Public Domain eBooks
 - Todins Unit Exam Chem Answer Keg eBook Subscription Services
 - Todins Unit Exam Chem Answer Keg Budget-Friendly Options
6. Navigating Todins Unit Exam Chem Answer Keg eBook Formats
 - ePub, PDF, MOBI, and More
 - Todins Unit Exam Chem Answer Keg Compatibility with Devices
 - Todins Unit Exam Chem Answer Keg Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Todins Unit Exam Chem Answer Keg
 - Highlighting and Note-Taking Todins Unit Exam Chem Answer Keg
 - Interactive Elements Todins Unit Exam Chem Answer Keg
8. Staying Engaged with Todins Unit Exam Chem Answer Keg
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Todins Unit Exam Chem Answer Keg
9. Balancing eBooks and Physical Books Todins Unit Exam Chem Answer Keg
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Todins Unit Exam Chem Answer Keg
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Todins Unit Exam Chem Answer Keg
 - Setting Reading Goals Todins Unit Exam Chem Answer Keg
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Todins Unit Exam Chem Answer Keg
 - Fact-Checking eBook Content of Todins Unit Exam Chem Answer Keg
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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