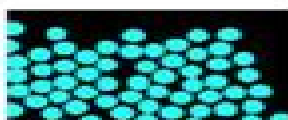


PhET Interactive Simulation: States of Matter

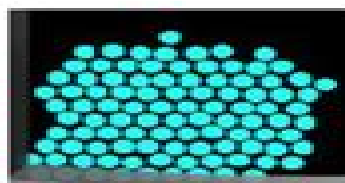
[Click Here for the States of Matter Simulation](#), practice then answer the questions below:
(Be sure to download the simulation before starting).

Exploration

1. Identify the following as either a solid, liquid, or a gas.



a) liquid



b) solid



c) gas

2. a) What happens to the atoms of a solid when the temperature increases?
b) decreases? *When the temp increases in a solid the atoms start to spread out, when the temp decreases the atoms get more and more compact.*
3. What happens to the atoms of a solid when the temperature decreases?
The atoms get more and more compact.
4. Complete the following table for the initial temperature of each substance in each state:

	Solid	Liquid	Gas
Neon	-295 c	-246 c	-217 c
Argon	-230 c	-187 c	-84 c
Oxygen	-246 c	-216 c	-160 c
Water	-127 c	13 c	156 c

5. Use the table above to compare any two (2) of the gases. How do their initial temperatures compare at each physical state?
When neon gets very cold it becomes a gas as you can see from the chart, but when water gets very hot it becomes a gas.
6. a) What happens to the pressure if the temperature is increased?
b) decreased? *When the temp is increased the pressure goes up for everyone of them except oxygen and when the temp is decreased the pressure goes down for everyone except oxygen. For oxygen when the temp is increased pressure goes down while decreasing it makes the pressure go up.*

States Of Matter Phet Answers

VM Jensen



States Of Matter Phet Answers:

This book delves into States Of Matter Phet Answers. States Of Matter Phet Answers is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into States Of Matter Phet Answers, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 3: States Of Matter Phet Answers in Everyday Life
 - Chapter 4: States Of Matter Phet Answers in Specific Contexts
 - Chapter 5: Conclusion
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