

States of matter

The three states of matter are solid, liquid and gas.

Solids:

- have a **fixed shape** and **cannot flow**, because their **particles cannot move** from place to place
- **cannot be compressed** (squashed), because their particles are close together and have no space to move into

Gases:

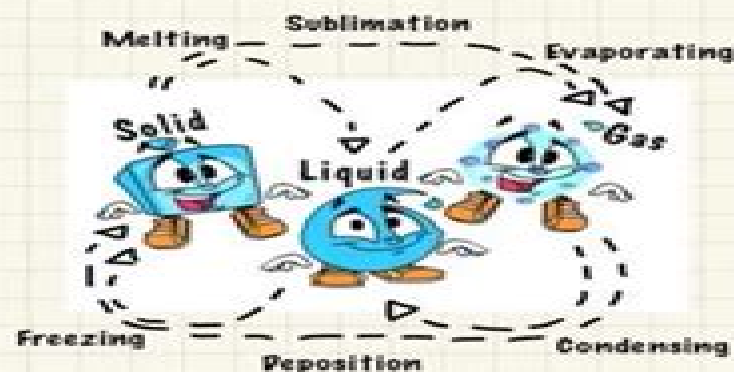
- flow and **completely fill their container**, because their **particles can move quickly** in all directions
- **can be compressed**, because their particles are far apart and have space to move into

Every substance has its own **melting point** and **boiling point**. The **stronger** the forces between particles, the **higher** its melting and boiling points.

Temperature	Given temperature < melting point
Predicted state	Solid
Temperature	Given temperature is between melting and boiling points
Predicted state	Liquid
Temperature	Given temperature > boiling point
Predicted state	Gas

Liquids:

- **flow** and take the **shape of their container**, because their **particles can move around** each other
- **cannot be compressed**, because their particles are close together and have no space to move into



Predicting a physical state

The state of a **substance** at a given temperature can be predicted if its melting point and boiling point are known. The table **summarises** how to work this out.

The particles in the states could be **atoms, molecules or ions** depending on the type of substance, eg **ionic compounds, small molecules, giant molecules and metals**.

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William Herbert Page



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