

Atomic Theory Worksheet

Name: _____

John Dalton (1766 – 1844):

John Dalton was an English chemist. His ideas form the atomic theory of matter. Here are his ideas.

Atomic Theory



- All elements are composed (made up) of atoms. It is impossible to divide or destroy an atom.
- All atoms of the same elements are alike. (One atom of oxygen is like another atom of oxygen.)
- Atoms of different elements are different. (An atom of oxygen is different from an atom of hydrogen.)
- Atoms of different elements combine to form a compound. These atoms must be in definite whole number ratios. For example, water is a compound made up of 2 atoms of hydrogen and 1 atom of oxygen (a ratio of 2:1). Three atoms of hydrogen and 2 atoms of oxygen cannot combine to make water.



1803

1. What is the name of John Dalton's theory? _____
2. What are elements made of? _____
3. An atom of hydrogen and an atom of carbon are _____.
4. What are compounds made of? _____
5. The ratio of atoms in HCl is: a) 1:3 b) 2:1 c) 1:1

J. J. Thompson (Late 1800s):

J. J. Thompson was an English scientist. He discovered the electron when he was experimenting with gas discharge tubes. He noticed a movement in a tube. He called the movement cathode rays. The rays moved from the negative end of the tube to the positive end. He realized that the rays were made of negatively charged particles – electrons.

Atomic Structure Model



J. J. Thompson



1904

1. What did J.J. Thompson discover? _____
2. What is the charge of an electron? _____
3. What are cathode rays made of? _____
4. Why do electrons move from the negative end of the tube to the positive end?

5. What was Thompson working with when he discovered the cathode rays?

NUCLEAR MODEL



ERNEST RUTHERFORD



1911

Lord Ernest Rutherford (1871 – 1937):

Ernest Rutherford conducted a famous experiment called the gold foil experiment. He used a thin sheet of gold foil. He also used special equipment to shoot alpha particles (positively charged particles) at the gold foil. Most particles passed straight through the foil like the foil was not there. Some particles went straight back or were deflected (went in another direction) as if they had hit something. The experiment shows:

- Atoms are made of a small positive nucleus; positive nucleus repels (pushes away) positive alpha particles
- Atoms are mostly empty space

1. What is the charge of an alpha particle? _____
2. How did he know that an atom was mostly empty space?

3. What happened to the alpha particles as they hit the gold foil?

4. How did he know that the nucleus was positively charged?

Student 1 Answers Atomic Theory

S Ashworth



Student 1 Answers Atomic Theory:

Student's Guide to Fundamentals of Chemistry Jo A. Beran, 2013-10-22 Student's Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles. This book deals with various approaches to chemical principles and problem solving in chemistry. Organized into 25 chapters, this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry. This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements. Other chapters consider how to calculate the molecular weight of a compound from its formula. This book discusses as well the characteristics of a photon in terms of its particle-like properties and defines the wavelength, frequency, and speed of light. The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters. This book is a valuable resource for chemistry students, lecturers, and instructors.

Physics for Students of Science and Engineering A. L. Stanford, J. M. Tanner, 2014-06-28 *Physics for Students of Science and Engineering* is a calculus-based textbook of introductory physics. The book reviews standards and nomenclature such as units, vectors, and particle kinetics, including rectilinear motion, motion in a plane, relative motion. The text also explains particle dynamics, Newton's three laws, weight, mass, and the application of Newton's laws. The text reviews the principle of conservation of energy, the conservative forces, momentum, the nonconservative forces, friction, and the fundamental quantities of momentum, mass, and velocity. The book examines changes in momentum, known as impulse, as well as the laws in momentum conservation in relation to explosions, collisions, or other interactions within systems involving more than one particle. The book considers the mechanics of fluids, particularly fluid statics, fluid dynamics, the characteristics of fluid flow, and applications of fluid mechanics. The text also reviews the wave-particle duality, the uncertainty principle, the probabilistic interpretation of microscopic particles such as electrons, and quantum theory. The book is an ideal source of reference for students and professors of physics, calculus, or related courses in science or engineering.

Colorado School Journal, 1899 *Prentice Hall Chemistry*, 2000 Contributions from Science Education Research Roser Pintó, Digna Couso, 2007-09-19 In August 2005, over 500 international researchers from the field of science education met at the 5th European Science Education Research Association conference in Barcelona, Spain. Two of the main topics at this conference were the decrease in the number of students interested in school science and concern about the worldwide outcomes of studies on students' scientific literacy. At the conference, over 400 papers were presented, covering a wide range of topics relevant to science education research, such as evidence-based practice, teachers' professional development, the role of ICT and multimedia, formal and informal learning environments, and argumentation and modelling in science education. This volume includes edited versions of 37 outstanding papers presented during the conference, including the lectures of the keynote speakers. They have been selected for their quality, variety, and interest and present a good overview of the field of science education research.

CLEP Success Peterson's, 2007-05-04 Offers advice on

improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers Physical Chemistry and Biophysics for Students of Biology and Medicine Matthew Steel,1928 *The Pharmaceutical Era* ,1895 **The Pharmaceutical Era** Charles W. Parsons,1920 **The Pharmaceutical Journal and Pharmacist** ,1909 *A Student's History of Philosophy* Arthur Kenyon Rogers,1907 *Pharmaceutical Journal* ,1897 *Pacific Pharmacist* W. M. Searby,Albert Schneider,1919 **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Proceedings of the Institution of Electrical Engineers Institution of Electrical Engineers,1924 Vols for 1970 79 include an annual special issue called IEE reviews Report of the ... Meeting of the British Association for the Advancement of Science British Association for the Advancement of Science,1888 **Saturday Review of Politics, Literature, Science and Art** ,1884 **Educational Review** ,1908 **Educational Review** Nicholas Murray Butler, Frank Pierrepont Graves,1908 Vols 19 34 include Bibliography of education for 1899 1906 compiled by James I Wyer and others Journal of Chemical Education ,1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

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