

Straighterline Chemistry Final- deck 3

Which is more soluble in an acidic solution than in pure water?

Select one:

- a. CuI
- b. PbCl_2
- c. $\text{Ca}_3(\text{PO}_4)_2$
- d. NaNO_3
- e. NaBr - ANSc. $\text{Ca}_3(\text{PO}_4)_2$

What is the name of the principle of selective precipitation used to identify the types of ions present in a solution?

Select one:

- a. Ionization
- b. Selective ion precipitation
- c. Selective ion typing
- d. Limited precipitation
- e. Qualitative analysis - ANSe. Qualitative analysis

A 10.0-mL sample of 0.75 M $\text{CH}_3\text{CH}_2\text{COOH}$ is titrated with 0.30 M NaOH . What is the pH of the solution after 22.0 mL of NaOH have been added to the acid?

$[\text{Ka}(\text{CH}_3\text{CH}_2\text{COOH}) = 1.3 \times 10^{-5}]$

Select one:

- a. 5.75
- b. 4.94
- c. 4.83
- d. 4.02
- e. 3.95 - ANSa. 5.75

Methyl red is a common acid-base indicator. It has a Ka equal to 6.3×10^{-6} . Its un-ionized form is red and its anionic form is yellow. What color would a methyl red solution have at $\text{pH} = 7.8$?

Select one:

- a. green
- b. red
- c. blue
- d. yellow
- e. violet - ANSd. yellow

A 25.0-mL sample of 0.10 M $\text{C}_2\text{H}_5\text{NH}_2$ (ethylamine) is titrated with 0.15 M HCl . What is the pH of the solution after 9.00 mL of acid have been added to the amine?

$[\text{Kb}(\text{C}_2\text{H}_5\text{NH}_2) = 6.5 \times 10^{-4}]$

Select one:

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introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students

RRB Junior Engineer (2019) - General Chemistry for CBT-1 & CBT-2 OnlineVerdan,2019-04-16 This Book encompasses all topics of General Chemistry according to syllabus of CBT 1 and CBT 2 of RRB Junior Engineer 2019 Exam The whole syllabus of General Chemistry is divided in eight sections First section of basics cover topic related to classification of Chemistry matter and its composition structure of atom periodic table chemical reactions unit systems etc Second and third sections describes about bonding between oxygen and hydrogen and carbon and nitrogen respectively Fourth section describes properties of various metals properties of acids and bases importance of nanotechnology in today s scenario Fifth section discusses about wider use of chemistry in agriculture food and medical sectors Sixth section is dedicated to polymers and its various varieties available in market Seventh section is related to chemical composition of fats and proteins The last eighth section detailed out thermodynamics and gas laws given by several scientists Further each section is divided in sub sections consisting detailed theory and practice questions The level of questions are easy to tough so that students may prepare not only for this exam but also other competitive exams such as UPSC CSAT State PSCs SSC JE etc The team OnlineVerdan have shown their best efforts to bring this unique book on e publication platform

General Chemistry Workbook Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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