

16. Increasing the temperature generally.....
- decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
 - increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
 - increases the solubility of both a solid and a gas in a liquid.
 - decreases the solubility of both a solid and a gas in a liquid.
17. A white precipitate was formed when $\text{BaCl}_{2(aq)}$ was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid effervescence. Salt X is likely to contain
- SO_4^{2-} ions.
 - SO_3^{2-} ions.
 - NO_3^- ions.
 - CO_3^{2-} ions.
18. Before a reaction could take place, there should be
- ionization of reactant particles.
 - breakage of bonds of products.
 - breakage of bonds of reactants.
 - ionization of product particles.
19. Consider the following reaction equation:
 $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$
 Silicon(IV) oxide is acting as
- a basic oxide.
 - a reducing agent.
 - an acidic oxide.
 - an oxidizing agent.
20. Which of the following acids would form normal salt only?
- Tetraoxosulphate(VI) acid
 - Trioxonitrate(V) acid
 - Tetraoxophosphate(V) acid
 - Trioxosulphate(IV) acid
21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm^3 of oxygen and 900 cm^3 of nitrogen gas?
- 1.0 atm
 - 0.1 atm
 - 9.0 atm
 - 0.9 atm
22. Graphite is used as a dry lubricant due to the presence of
- mobile electrons.
 - free electrons.
 - octahedral structures.
 - layered structures.
23. Which of the following gases is alkaline?
- NO_2
 - CO_2
 - NH_3
 - N_2O
24. Which of the following statements about an equilibrium system is correct?
- Forward and backward reactions occur at the same rate.
 - The concentrations of reactants must equal that of the products.
 - The concentrations of reactants and products can be changed by adding a catalyst.
 - The concentrations of reactants and products are not affected by a change in temperature.

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 - The Rise of Digital Reading Waec Chemistry Question Paper 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Waec Chemistry Question Paper 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Waec Chemistry Question Paper 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Waec Chemistry Question Paper 2014
 - Personalized Recommendations
 - Waec Chemistry Question Paper 2014 User Reviews and Ratings
 - Waec Chemistry Question Paper 2014 and Bestseller Lists
5. Accessing Waec Chemistry Question Paper 2014 Free and Paid eBooks

- Waec Chemistry Question Paper 2014 Public Domain eBooks
- Waec Chemistry Question Paper 2014 eBook Subscription Services
- Waec Chemistry Question Paper 2014 Budget-Friendly Options
- 6. Navigating Waec Chemistry Question Paper 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Waec Chemistry Question Paper 2014 Compatibility with Devices
 - Waec Chemistry Question Paper 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Waec Chemistry Question Paper 2014
 - Highlighting and Note-Taking Waec Chemistry Question Paper 2014
 - Interactive Elements Waec Chemistry Question Paper 2014
- 8. Staying Engaged with Waec Chemistry Question Paper 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Waec Chemistry Question Paper 2014
- 9. Balancing eBooks and Physical Books Waec Chemistry Question Paper 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Waec Chemistry Question Paper 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Waec Chemistry Question Paper 2014
 - Setting Reading Goals Waec Chemistry Question Paper 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Waec Chemistry Question Paper 2014
 - Fact-Checking eBook Content of Waec Chemistry Question Paper 2014
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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