

16. Increasing the temperature generally.....
- decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
  - increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
  - increases the solubility of both a solid and a gas in a liquid.
  - decreases the solubility of both a solid and a gas in a liquid.
17. A white precipitate was formed when  $\text{BaCl}_{2(aq)}$  was added to an aqueous solution of a salt X. The precipitate dissolved in dilute  $\text{HCl}$  with rapid effervescence. Salt X is likely to contain
- $\text{SO}_4^{2-}$  ions.
  - $\text{SO}_3^{2-}$  ions.
  - $\text{NO}_3^-$  ions.
  - $\text{CO}_3^{2-}$  ions.
18. Before a reaction could take place, there should be
- ionization of reactant particles.
  - breakage of bonds of products.
  - breakage of bonds of reactants.
  - ionization of product particles.
19. Consider the following reaction equation:  
 $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$   
 Silicon(IV) oxide is acting as
- a basic oxide.
  - a reducing agent.
  - an acidic oxide.
  - an oxidizing agent.
20. Which of the following acids would form normal salt only?
- Tetraoxosulphate(VI) acid
  - Trioxonitrate(V) acid
  - Tetraoxophosphate(V) acid
  - Trioxosulphate(IV) acid
21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing  $100 \text{ cm}^3$  of oxygen and  $900 \text{ cm}^3$  of nitrogen gas?
- 1.0 atm
  - 0.1 atm
  - 9.0 atm
  - 0.9 atm
22. Graphite is used as a dry lubricant due to the presence of
- mobile electrons.
  - free electrons.
  - octahedral structures.
  - layered structures.
23. Which of the following gases is alkaline?
- $\text{NO}_2$
  - $\text{CO}_2$
  - $\text{NH}_3$
  - $\text{N}_2\text{O}$
24. Which of the following statements about an equilibrium system is correct?
- Forward and backward reactions occur at the same rate.
  - The concentrations of reactants must equal that of the products.
  - The concentrations of reactants and products can be changed by adding a catalyst.
  - The concentrations of reactants and products are not affected by a change in temperature.

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