## **AP Chemistry Final Study Guide**

#### Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- Mass Spectroscopy:
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- Coulomb's Law:
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- · Electromagnetic Radiation:
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- Kinetic Energy of Ionized Electrons:
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so
    it will have more kinetic energy
- · Electron Configuration:
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel: 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>3d<sup>8</sup>4s<sup>2</sup>(notice the d-block)
    - Nickel (with shorthand): [Ar]4s<sup>2</sup>3d<sup>8</sup>
- Law of Definite Proportions:
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- Law of Multiple Proportions:
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- Aufbau Principle:
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# **Spring Exam Study Guide Chemistry**

Peter L. Kresan, Reed Mencke, Frank Press, Raymond Siever

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come to school with innate curiosity and intuitive ideas about the world around them science classes rarely tap this potential Many experts have called for a new approach to science education based on recent and ongoing research on teaching and learning In this approach simulations and games could play a significant role by addressing many goals and mechanisms for learning science the motivation to learn science conceptual understanding science process skills understanding of the nature of science scientific discourse and argumentation and identification with science and science learning To explore this potential Learning Science Computer Games Simulations and Education reviews the available research on learning science through interaction with digital simulations and games It considers the potential of digital games and simulations to contribute to learning science in schools in informal out of school settings and everyday life The book also identifies the areas in which more research and research based development is needed to fully capitalize on this potential Learning Science will guide academic researchers developers publishers and entrepreneurs from the digital simulation and gaming community and education practitioners and policy makers toward the formation of research and development partnerships that will facilitate rich intellectual collaboration Industry government agencies and foundations will play a significant role through start up and ongoing support to ensure that digital games and simulations will not only excite and entertain but also motivate and educate

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