

DATE: 10/20/21  
 SECTION/GROUP: #10  
 NAME: Bailey Spell  
 NAME: Gavin Sprinkle  
 NAME: Ben Steadley

### Investigation 3 Written Argument

Guiding Question: What is the concentration of acetic acid in vinegar?

Claim: My group and I claim that the concentration of acetic acid in vinegar is 0.83 M, 4.99% m/v

Evidence: Balanced equation:  $1\text{CH}_3\text{COOH}(\text{aq}) + 1\text{NaOH}(\text{aq}) \gg> 1\text{NaCH}_3\text{COO}(\text{aq}) + 1\text{H}_2\text{O}(\text{l})$

#### Data Table

| Data:                          | Trial 1:   | Trial 2:   | Trial 3:  | Trial 4:  |
|--------------------------------|------------|------------|-----------|-----------|
| Molarity of NaOH used          | 0.400 M    | 0.400 M    | 0.400 M   | 0.400 M   |
| Volume of Acetic Acid used     | 10.00 mL   | 10.00 mL   | 10.00 mL  | 10.00 mL  |
| Initial Burette Reading (NaOH) | 0.35 mL    | 21.15 mL   | 0.49 mL   | 21.21 mL  |
| Final Burette Reading (NaOH)   | 21.15 mL   | 41.87 mL   | 21.21 mL  | 42.15 mL  |
| Volume of NaOH used            | 20.80 mL   | 20.72 mL   | 20.72 mL  | 20.94 mL  |
| Shade of Pink                  | Light pink | Light pink | Dark pink | Dark pink |

#### Calculations of volume of NaOH used:

**Trial 1: Initial Reading - Final Reading**

21.15 mL NaOH - 0.35 mL of NaOH = 20.80 mL NaOH

**Trial 2: Initial Reading - Final Reading**

41.87 mL NaOH - 21.15 mL of NaOH = 20.72 mL NaOH

**Trial 3: Initial Reading - Final Reading**

21.21 mL NaOH - 0.49 mL NaOH = 20.72 mL NaOH

**Trial 4: Initial Reading - Final Reading**

42.15 mL NaOH - 21.21 mL NaOH = 20.94 mL NaOH

#### Calculations of moles of NaOH used:

Moles of NaOH used = Molarity x Volume in L

\*Make sure to change mL of NaOH into L of NaOH\*

**Trial 1:**  $20.80 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02080 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02080 \text{ L NaOH} = 0.00832 \text{ mols NaOH}$

**Trial 2:**  $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

**Trial 3:**  $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

**Trial 4:**  $20.94 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02094 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02094 \text{ L NaOH} = 0.00838 \text{ mols NaOH}$

**Calculations of moles and Molarity of Acetic Acid:**

mols of Acetic Acid is same as mols of NaOH due to balanced equation

\*Make sure to change mL of volume of acetic acid used to L of acetic acid used\*  $\gg> \gg$

$(10.00 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}}) / 1000 \text{ mL} = 0.010 \text{ L}$

Molarity = mols/L

**Trial 1:**  $0.00832 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

**Trial 2:**  $0.00829 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

# Us Ws 1 V3 0 Answers

**John D. Cutnell, Kenneth W.  
Johnson, David Young, Shane Stadler**

## Us Ws 1 V3 0 Answers:

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