

Preparation and standardization of solution AgNO_3 0.02 mol/L

1. Introduction

The precipitation volumetric is based on reactions with formation of poorly soluble compounds. The precipitation reaction must be almost quantitatively carried out at the equivalence point, complete in relatively short time and provide conditions for convenient end point signaling.

In practice, such conditions greatly limit the number of precipitation reactions that can be used. Many of them cannot serve because of the lack of appropriate means for locating the endpoint. In a small number of cases, it is possible to conduct the titration under visual observation to the point where precipitate formation ceases to occur. Most commonly, the use of indicators is called for.

Many volumetric methods of precipitation employ more or less specific indicators, i.e. appropriate for a given precipitation reaction. There is, however, a special class of indicators, the **adsorption indicators**, which find a more general field of application. The possibilities of applying precipitation reactions in volumetric analysis are greatly expanded with the use of physico-chemical methods for the location of the end point.

The adsorption indicators were introduced by Fajans. They are organic dyes, with the character of weak acids or bases (anionic or cationic, respectively), which accuse the end point through a change of coloration on the precipitate. The change of color is due to adsorption or desorption of the dye as a consequence of a modification of the electric double layer around the precipitate particles at the point of equivalence: thus, the appearance or disappearance of a coloration on the precipitate serves to signal the end point.

In volumetric analyzes in general, the variation of the concentrations of the reagent ions around the point of equivalence depends greatly on the degree to which the reaction completes. In precipitation volumetry, the factors that decide the question are the product of the solubility of the precipitate and the concentrations of the reagents.

The most important volumetric method of precipitation, unique with a more or less wide field of application, is argentimetry, which is based on the formation of poorly soluble silver salts (halides, cyanide, thiocyanate). There are basically three argentimetric methods: Mohr method, Volhard method and by adsorption indicators.

Argentimetry involves the use of standard solutions of silver nitrate (AgNO_3) and has as its main field of application to the determination of halides and other anions that form poorly soluble silver salts. Argentimetry comprises different methods, which can be classified according to whether the titration is direct or indirect.

In direct methods, the solution containing the substance to be determined is titrated with silver nitrate standard solution to the point of equivalence. The endpoint can be identified in several ways: addition of silver nitrate until no more is observed the formation of precipitate or change of color of an indicator.

Standardization Of Agno3 Lab Report

Sebastian Brünink

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Laboratory Manual for Chemistry: a Quantitative Approach Robert Nelson Smith, 1969 Chemical Abstracts ,1910
Analytical Chemistry Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry **Analytical Chemistry** Mr. Rohit Manglik, 2024-07-03 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels **Biological Abstracts** ,1928 Agrindex ,1991
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