

Unit 10 Stoichiometry Test 1: Short Answers

Answer all questions including units and balanced equations where appropriate.

1. How many moles of aluminium are produced when completely with 2.5 mol of FeCl_3 ?

$$2\text{Al} + 3\text{FeCl}_2 \rightarrow 3\text{Fe} + 2\text{AlCl}_3$$
2. How many grams of FeCl_3 are produced when 2.5 moles of aluminium with oxygen to FeCl_3 ?

$$2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$$
3. The reaction below shows that decomposition of FeCl_3 to FeCl_2 and Cl_2 . How many grams of oxygen are produced when 20.7g FeCl_3 is heated?

$$2\text{FeCl}_3 \rightarrow 2\text{FeCl}_2 + \text{Cl}_2$$
4. A reaction that has been calculated to produce 2.0g of FeCl_3 actually produces 2.2g of FeCl_3 . What is the percent yield?
5. How many moles of FeCl_3 are produced when 20.7g of FeCl_3 is heated?

$$2\text{FeCl}_3 \rightarrow 2\text{FeCl}_2 + \text{Cl}_2$$

6. Percent Yield is

7. Amount of 20.7g produced is

8. How many moles of FeCl_3 are produced when 20.7g of FeCl_3 is heated?

Unit 8 Test Stoichiometry

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Luo, Xiaodong Zhang**

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