

UNIT 5: CHEMICAL KINETICS AND EQUILIBRIUM

Chapter 14: Chemical Kinetics

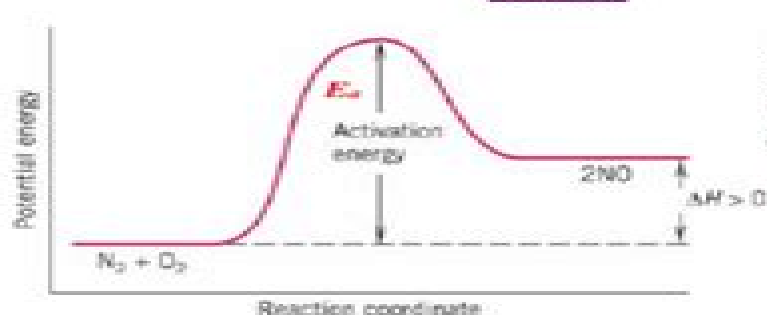
14.4 & 14.6: Activation Energy, Temperature Dependence on Reaction Rates & Catalysis

Reaction Rates - the speed of which the concentration of a reactant or product changes over time.

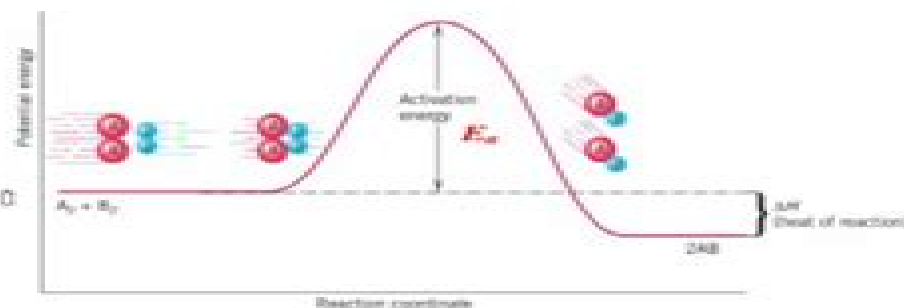
Collision Model - a model that state for a reaction to occur, molecules must collide with each other.

Factors Affecting the Collision Model

1. **Activation Energy (E_a)** - the threshold energy molecules needed to overcome to cause a chemical reaction that was first proposed by Svante Arrhenius.
 - E_a is the highest energy (top of the hill - E_{peak}) minus the sum of energy of the reactants ($\Sigma H_{\text{reactants}}$) on the potential energy diagram.
 - in general, the **Lower the Activation Energy, the Faster is the Rate of Reaction**

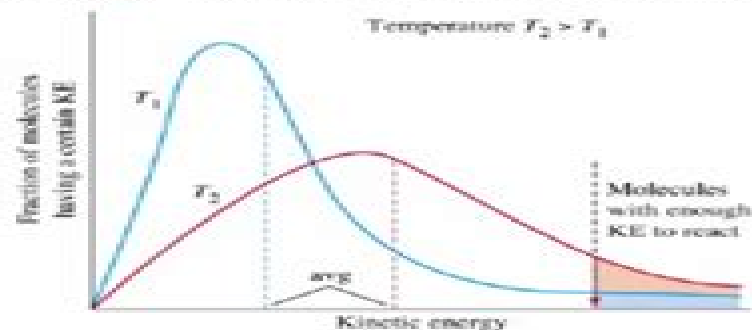


Activation Energy of an Endothermic Reaction



Activation Energy of an Exothermic Reaction

Temperature (T) - the effective number of collisions increases exponentially with temperature.



Because $T_2 > T_1$, there are more molecules colliding with enough Kinetic Energy causing a reaction to occur.

In general, the **Higher the Temperature, the Faster the Rate of Reaction.**

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