



Empirical and Molecular Formulas Worksheet

1. The molecular formulas of some substances are as follows. Write their empirical formulas.

a. Acetylene, C_2H_2 (used in oxyacetylene torches) _____

b. Glucose, $C_6H_{12}O_6$ (the chief sugar in blood) _____

c. Octane, C_8H_{18} (a component of gasoline) _____

d. Ethylene glycol is $C_2H_6O_2$ (antifreeze) _____

2. Determine the empirical formula of a compound that contains 36.5% sodium, 25.4% sulfur, and 38.1% oxygen.

3. Nitrogen and oxygen form an extensive series of oxides with the general formula N_xO_y . One of them is a blue solid that comes apart, reversibly, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide?

4. An organic compound has an empirical formula of CH and a molecular mass of 78 g/mol. What is the molecular formula?

5. A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C_2H_4NO . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound?

6. Determine the molecular formula for each compound: (remember that you need to find empirical formula first!)

a. 94.1% O and 5.9% H; molar mass = 34 g

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