

Answer KEY

Unit 3 Outline/ Study Guide - Electron Configuration and Periodic Trends

Essential Skills/ State Standards:

1. Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, halogens, alkali metals, alkaline earth metals, and transition metals.
3. Students know how to use the periodic table to identify the following trends:
ionization energy, electronegativity, and the relative sizes of ions and atoms
(this includes knowing the definitions of ionization energy, electronegativity, cation, and anion.)

Study Guide - (use your notes and worksheets for a more complete review)

1. Identify the following things on the periodic table:

-metals	-alkali metals
-metalloids (semi-metals)	-alkaline earth metals
-transition metals	-halogens
-nonmetals	-noble gases
-Periods	-Groups/ families
Identify groups with the following numbers of valence electrons: 1, 2, 3, 4, 5, 6, 7, 8	

2. Why do the elements in the groups (columns) behave so similar to each other even if they have great differences in their number of protons? **same number of valence electrons and similar chemical properties.**
3. What is the connection between the following terms: **octet rule, valence electrons, and ions?**
In order to achieve the octet rule, atoms gain or lose their valence electrons. They become charged ions when this occurs.
4. a. List the elements that are exceptions to the octet rule b) Why don't these atoms achieve the octet rule?
a. Small atoms are happy with 2 electrons (ex: H, He, Li, Be, B)
b. because the 1s energy level only holds 2 electrons

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A Gutmann



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Table of Contents Unit 3 Periodic Table Test Answer

1. Understanding the eBook Unit 3 Periodic Table Test Answer
 - The Rise of Digital Reading Unit 3 Periodic Table Test Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 3 Periodic Table Test Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 3 Periodic Table Test Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 3 Periodic Table Test Answer
 - Personalized Recommendations
 - Unit 3 Periodic Table Test Answer User Reviews and Ratings
 - Unit 3 Periodic Table Test Answer and Bestseller Lists
5. Accessing Unit 3 Periodic Table Test Answer Free and Paid eBooks
 - Unit 3 Periodic Table Test Answer Public Domain eBooks

- Unit 3 Periodic Table Test Answer eBook Subscription Services
- Unit 3 Periodic Table Test Answer Budget-Friendly Options
- 6. Navigating Unit 3 Periodic Table Test Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 3 Periodic Table Test Answer Compatibility with Devices
 - Unit 3 Periodic Table Test Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 3 Periodic Table Test Answer
 - Highlighting and Note-Taking Unit 3 Periodic Table Test Answer
 - Interactive Elements Unit 3 Periodic Table Test Answer
- 8. Staying Engaged with Unit 3 Periodic Table Test Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 3 Periodic Table Test Answer
- 9. Balancing eBooks and Physical Books Unit 3 Periodic Table Test Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 3 Periodic Table Test Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 3 Periodic Table Test Answer
 - Setting Reading Goals Unit 3 Periodic Table Test Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 3 Periodic Table Test Answer
 - Fact-Checking eBook Content of Unit 3 Periodic Table Test Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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