

Unit 3 - Periodicity

Trends in physical properties

Effective Nuclear Charge

- The nuclear charge of an atom is given by its atomic number. (e.g. Oxygen is +8)
- The outer electrons of the atom **do not experience the full attraction** of the nuclear charge because they are **shielded from the nucleus because of the repulsion by the inner electrons** and so the effective nuclear charge is lesser than the nuclear charge.
- Nuclear charge increases across a period, but remains constant down a group

Atomic Radius

- the distance from the nucleus to the outermost electrons of an atom.
- Atomic radius decreases across a period, as attraction between the nucleus and outermost electrons increases, and increases down a group, as the number of shells increases

Ionic Radius

- Cations: As there are less electrons than protons, the outer shell is lost and the ionic radius of a cation is lesser than the atomic radius of the atom.
 - e.g. Na has a atomic radius of 186×10^{-12} while Na^+ has a ionic radius of 98×10^{-12}
- Anions: As there are more electrons than protons, there is more electron repulsion and less attraction, so the ionic radius is more than the atomic radius of the atom.
 - e.g. Cl has an atomic radius of 99×10^{-12} , but Cl^- has an ionic radius of 181×10^{-12}

Ionization Energy

- The first ionization energy of an element is the energy required to remove one mole of electrons from one mole of gaseous atoms in their ground state.
- Increases across a period as atomic bonds are stronger, and decreases down a group as there are more outshells and attraction is weaker.

Exceptions

- Groups 2 and 3
 - Elements of group 3 have a lower lower ionization energy than elements of Group 2.
 - Elements of Group 3 have the electron configuration of s^2p^1 , which means that the electron removed is from the p orbital, which is unpaired, making it easier to remove.
 - However, elements of Group 2 have an electron configuration of s^2 , which is a complete orbital and when an electron is removed, it is removed from this orbital, and it is relatively harder to remove the electron.
- Groups 15 and 16
 - Elements of Group 16 have a lower ionization energy compared to elements of Group 15.

Unit 3 Electron Arrangement And Periodicity Study Guide Answers

Sasha Hu



Unit 3 Electron Arrangement And Periodicity Study Guide Answers:

Unveiling the Magic of Words: A Report on "**Unit 3 Electron Arrangement And Periodicity Study Guide Answers**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Unit 3 Electron Arrangement And Periodicity Study Guide Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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