

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the flask solution. This would be incorrect. **Just this solution.**
Then, calculate the molarities of the two solutions. The molarity of the solution is 1.5M. The molarity of the solution is 1.5M.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} = \frac{1 \text{ mol}}{74.55 \text{ g/mol}} \times 100.0 \text{ mL}$$

$$10.0 \text{ g} = \frac{1 \text{ mol}}{74.55 \text{ g/mol}} \times 100.0 \text{ mL}$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$25.0 \text{ g} = \frac{1 \text{ mol}}{110.98 \text{ g/mol}} \times 0.5 \text{ L}$$

$$25.0 \text{ g} = \frac{1 \text{ mol}}{110.98 \text{ g/mol}} \times 0.5 \text{ L}$$

4. An unknown liquid solution of "molar" and "molar" is prepared by dissolving 5.00 g of sodium chloride in enough water to make a solution of 100.0 mL. Calculate the molarity of sodium in this solution.

$$5.00 \text{ g} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL}$$

$$5.00 \text{ g} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL}$$

5. A 10.0 g solution of sodium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molarity of sodium in this solution.

$$10.0 \text{ g} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL}$$

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VM Jensen



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Nuclear Technology Programs Semiannual Progress Report, October 1989-March 1990, 1992

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