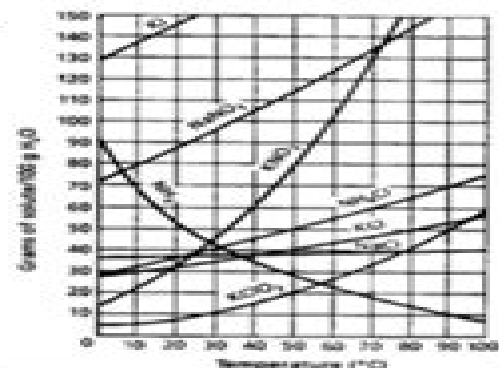


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SOLUTIONS WORKSHEET

A. Solubility curves – refer to the graph to answer the following questions.



1. What is the solubility of NH_4Cl at 70°C ? **60 g**
2. Calculate the mass of NH_4Cl that would dissolve in 200 g of water at 90°C .

70 g @ 100g water $\rightarrow$ 140 g

3. At 0°C , 50.0 g NH_4Cl are added to 100 g water. To what temperature does the solution have to be raised in order to achieve a saturated solution of $\text{NH}_4\text{Cl}_{\text{sat}}$?

50°C

4. Is NH_3 a (s), (l), or (g)? Explain.

GAS – it's solubility decreases with increase in temperature. Opposite for solids

5. Determine how much water is required to dissolve 35.0 g NH_4Cl at 70°C in order to make a saturated solution. Show your work.

$$\frac{60}{100} = \frac{35}{x}$$

$$x = 58.3$$

$$x = 58.3 \text{ g H}_2\text{O}$$

6. If a saturated solution of NH_3 at 15°C is warmed to 40°C , how much ammonia will "fall out" of the solution?

@ 15°C – 60 g soluble

@ 40°C – 35 g soluble

25 g will fall out

B. MOLARITY

1. What is molarity a measure of? What is the formula for molarity? What are the units of molarity?

Concentration.

$$M = n / V$$

mol/L or M

2. Calculate the molarity of a 1.5 L solution that has 0.45 mol solute dissolved.

0.30 - M

3. Calculate the grams of sodium sulfate that are required to prepare 1.50 L of 0.25 M solution.

53.3 g Na_2SO_4

4. Calculate the volume of solution required to make 0.024 M solution using 15.0 g of calcium chloride.

V = 5.6 L

5. Describe in detail the steps you would take to prepare 1.50 L of 0.5 M solution of sodium chloride.

$$n = 0.75 \text{ mol} \rightarrow 43.5 \text{ g NaCl}$$

1) measure out 43.5 g NaCl

2) put NaCl into graduated cylinder

3) add enough water to make final volume 1.50 L

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