

# Build an Atom Virtual Lab

3 pages + 6-page Answer Key

Chemistry

Physical Science

## BUILD AN ATOM VIRTUAL LAB

### Introduction:

In this lab you will understand the structure of atoms, including protons, neutrons, and electrons, and how these particles determine the identity and properties of elements.

### Materials:

Virtual Lab: Build an Atom (https://phet.colorado.edu/en/simulations/build-an-atom)

### Procedure (Part 1)

1. Access the Lab: Enter type in the search bar.
2. Build at least 8 different atoms by adjusting the number of protons, neutrons, and electrons. Record the atomic number, mass number, and charge of each atom in the table below.
3. Observe the effect of adding or removing a proton, neutron, or electron on the atom's identity and properties.

| Element | Symbol | Protons | Neutrons | Electrons | Charge |
|---------|--------|---------|----------|-----------|--------|
|         |        |         |          |           |        |
|         |        |         |          |           |        |
|         |        |         |          |           |        |
|         |        |         |          |           |        |

Atoms



UNDER CONSTRUCTION

When building an atom, write the information:

1. Atomic Number (Z): Number of protons in the nucleus.

2. Mass Number (A): Sum of protons and neutrons in the nucleus.

3. Charge (Q): Difference between the number of protons and electrons.

4. Element Symbol: Chemical symbol for the element.

5. Isotope Name: Name of the specific isotope.

6. Atomic Weight: Average mass of the atom.

7. Periodic Table: Location of the element.

8. Group Number: Number of the group in the periodic table.

9. Period Number: Number of the period in the periodic table.

10. Block: s-block, p-block, d-block, or f-block.

11. Valence Electrons: Number of electrons in the outermost shell.

12. Ionization Energy: Energy required to remove an electron.

13. Electronegativity: Ability of an atom to attract electrons.

14. Atomic Radius: Distance from the nucleus to the outermost shell.

15. Density: Mass per unit volume.

16. Melting Point: Temperature at which the solid phase changes to the liquid phase.

17. Boiling Point: Temperature at which the liquid phase changes to the gas phase.

18. Half-life: Time taken for half of the substance to decay.

19. Radioactivity: Emission of particles or energy from the nucleus.

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**D Keegan**



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