

Energy Storage and Transfer Model Worksheet Quantitative Energy Calculations & Energy Conservation

We use the table below each time we calculate:

1. When we calculate kinetic energy, mass is **2000 kg** (assuming 2000 kg car) or **500 kg** (assuming 500 kg car)

$$K = \frac{1}{2}mv^2 = \frac{1}{2}(2000 \text{ kg})v^2$$

$$K = \frac{1}{2}mv^2 = \frac{1}{2}(500 \text{ kg})v^2 \quad \text{if we assume we have a 500 kg car}$$

2. When we calculate energy stored in the **spring**, both quantities of **displacement** and **spring** (force constant) are given in **1000 N/m**

$$U_{\text{spring}} = \frac{1}{2}kx^2 = \frac{1}{2}(1000 \text{ N/m})x^2$$

3. Kinematic given: If **gravitational** forces acting on the object are the **only** forces, the **gravitational** energy of the object, which is equal to the state reference level is a **constant** value:

$$U_g = mgh$$

or, also called

$$U_g = mgy \quad \text{if } y = h$$

- 4a. When **force** is **constant**, before the third:

$$U_s = \frac{1}{2}kx^2 \quad \text{if } x = 0.20 \text{ m} \Rightarrow U_s = 20 \text{ J}$$

- or, when **velocity** is **constant** about the third:

$$U_s = \frac{1}{2}mv^2 \quad \text{if } v = 2.0 \text{ m/s} \Rightarrow U_s = 20 \text{ J}$$

$$U_s = \frac{1}{2}mv^2 \quad \text{if } v = 2.0 \text{ m/s} \Rightarrow U_s = 20 \text{ J}$$

- 4b. A **spring** with **displacement** (distance) and **spring constant** of **1000 N/m**. When **elastic** energy is **stored** (unbalanced):

$$U_s = \frac{1}{2}kx^2 = \frac{1}{2}(1000 \text{ N/m})(0.20 \text{ m})^2 = 20 \text{ J}$$

5. When there is a **spring** and **spring constant** given, the **potential** (elastic) energy is:

$$U_s = \frac{1}{2}kx^2 = \frac{1}{2}(1000 \text{ N/m})(0.20 \text{ m})^2 = 20 \text{ J}$$

6. When **total** energy is **constant** (if it's **isolated** and **spring** is **unbalanced** all the time)

→ When **spring** is **unbalanced** about **2000 N/m** (unbalanced) **2000 N/m**, $x = 0.20 \text{ m}$

$$U_s = \frac{1}{2}kx^2 = \frac{1}{2}(2000 \text{ N/m})(0.20 \text{ m})^2 = 40 \text{ J}$$

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The Resource Recovery Planning Model Robert E. Chapman, Edward B. Berman, 1983 *Sample Campus Environmental Audits*, 2005 Mn Manganese, 2013-06-29 The present volume Manganese D 5 continues the description of the manganese complexes The arrangement of the complexes in these D volumes is based on the ligand type The introduction on p 1 shows the classes of complexes which have already been described in Chapters 1 to 21 in the Volumes D 1 1979 D 2 1980 D 3 1982 and D 4 1985 In Chapters 22 to 29 of this volume are treated complexes with amine N polycarboxylic acids hydrazinecarboxylic acids amides hydrazides derivatives of hydroxylamine e g hydroxamic acids oximes and nitroso compounds azo compounds and triazenes A survey at the beginning of each of these sections gives information on the most characteristic features of the various complex types Because of the complexometric relevance of the complexes with amine N polycarboxylic acids there are many studies concerning the existence and the stability of the complexes in solution Numerous X ray investigations reported for the complexes with urea or with amides and hydrazides of carboxylic acids show the different structures of the compounds as a result of the varying bonding sites of the ligands Complexes with hydrazides e g with isonicotino hydrazide are of special interest due to their biological activity Complexes with hydroxamic acids oximes or azo compounds have been studied mostly in aqueous organic or pure organic solvents The characteristic intense colors of many solutions are used for the analytical determination of manganese Fe Organoiron Compounds Adolf Slawisch, 2013-06-29 *Cu Organocopper Compounds* Helmut Bauer, Jürgen Faust, Rolf Froböse, Johannes Füssel, Ulrich Krüerke, Manfred Kunz, Herman M. Somer, 2013-06-29 The present volume finalizes the coverage of organocopper compounds A complete formula and ligand index for the Gmelin organocopper series will appear shortly as *Organocopper Compounds 5* The volume describes mononuclear compounds with ligands bonded by two or more carbon atoms as well as all di to octanuclear and polymeric compounds Mononuclear compounds with ligands bonded by one carbon atom have already been described in Volumes 1 published in 1985 2 published in 1983 and 3 published in 1986 As structural elucidation in organocopper chemistry gained more attention only in the last few years the terms mononuclear dinuclear etc have been used as explained in *Organocopper Compounds 1* 1985 pp 3 4 all compounds are treated with their smallest formula unit unless a higher nuclearity has been proved As a consequence most of the species treated in volumes 1 to 3 are described there because of insufficient structural information although they are alleged not to be monomeric This way many of the better characterized compounds appear in the present volume which is reflected by the more than eighty X ray structure figures Generally nuclearity and structure are not only determined by the coordination properties of the ligands but also by steric requirements and may therefore widely differ for analogous compounds For abbreviations and dimensions used throughout this volume see p X Frankfurt am Main July 1987 Johannes Füssel Remarks on Abbreviations and Dimensions Most compounds and reagents in this volume are presented in tables **Energy Research Abstracts**, 1978 Semiannual

with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes **Sb Organoantimony Compounds Part 4** Markus Wieber,2013-06-29

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