

Unit 6: Stoichiometry Homework Packet

Name: _____

Period: _____

Mole Ratios:

1. What is a mole ratio?

2. The coefficients in a chemical equation indicate the relationship among _____ of reactants and products.

3. Write 3 possible mole ratios from this equation: $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

4. Write 4 possible mole ratios from this equation: $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

5. Write a balanced chemical equation when nitrogen reacts with hydrogen to produce nitrogen trihydride. Determine 3 possible mole ratios.

6. Stoichiometry is based on the _____.

7. This law states that _____.

Write the mole ratio in the space provided and solve each problem.



8. How many moles of sodium hydroxide are produced when 1.50 mol sodium peroxide reacts with water?

$$1.50 \text{ mol Na}_2\text{O}_2 \quad \times \quad \frac{4 \text{ mol NaOH}}{2 \text{ mol Na}_2\text{O}_2} \quad = \quad \underline{3.00} \quad \text{mol NaOH}$$

9. How many moles of water are required to produce 2.15 mol oxygen gas in this reaction?

$$2.15 \text{ mol O}_2 \quad \times \quad \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \quad = \quad \underline{4.30} \quad \text{mol H}_2\text{O}$$

Stoichiometry Practice Problems #1..... Solve each problem, show all work.

1. How many **grams** of Cl_2 must be reacted with excess NaI if 10.0g of NaCl are needed?



Unit 7 Homework 6 Molarity Stoichiometry

Liyong Dong

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Unit 7 Homework 6 Molarity Stoichiometry:

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