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Unit 3 Worksheet 4 – Quantitative Energy Problems Part 2

Energy constants (H2O)

334 J/g Heat of fusion (melting or freezing) H_f

2260 J/g Heat of vaporization (evaporating or condensing) H_V

2.1 J/g°C Heat capacity (c) of solid water 4.18 J/g°C Heat capacity (c) of liquid water

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

 How much energy must be absorbed by a 150 g sample of ice at 0.0 °C that melts and then warms to 25.0 °C?

Suppose in the ley Hot lab that the burner transfers 325 kJ of energy to 450 g of liquid water at 20. °C. What mass of the water would be boiled away?

3. A 12oz can of soft drink (assume m = 340 g) at 25°C is placed in a freezer where the temperature is - 12°C. How much energy must be removed from the soft drink for it to reach this temperature?

Unit Worksheet Quantitative Energy Problems

Sabine Zange

Unit Worksheet Quantitative Energy Problems:

This book delves into Unit Worksheet Quantitative Energy Problems. Unit Worksheet Quantitative Energy Problems is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Unit Worksheet Quantitative Energy Problems, encompassing both the fundamentals and more intricate discussions.

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